

- (c) Using the redox systems in **Table 6.1**, construct an equation for a reaction between acidified dichromate(VI) ions and methanoic acid, HCOOH .

Rather than using $[\text{O}]$ or $[\text{H}]$, your equation must show the actual reactants and products.

[2]

- (d) A student added some chromium metal to an acidified solution containing copper(II) ions. A reaction took place. The student concluded that 'chromium is more reactive than copper'.

- (i) Explain, in terms of their electrode potentials, why 'chromium is more reactive than copper' in this reaction.

.....

 [2]

- (ii) When this experiment was carried out, the student observed some bubbles of a gas. Suggest an explanation for this observation.

.....
 [1]

- (e) Methanoic acid, HCOOH , has the common name of 'formic acid'. Direct-Formic Acid Fuel Cells (DFAFCs) are being developed for use in small, portable electronics such as phones and laptop computers.

In this fuel cell, methanoic acid (the fuel) reacts with oxygen to generate a cell potential.

- (i) Predict the standard cell potential of a DFAFC.

standard cell potential = V [1]

- (ii) Suggest **two** advantages of using methanoic acid as the fuel in a fuel cell rather than hydrogen.

.....

 [2]

[Total: 14]

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1. Understanding the eBook Ocr F29june 2012 Question Paper
 - The Rise of Digital Reading Ocr F29june 2012 Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr F29june 2012 Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr F29june 2012 Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr F29june 2012 Question Paper
 - Personalized Recommendations
 - Ocr F29june 2012 Question Paper User Reviews and Ratings

- Ocr F29june 2012 Question Paper and Bestseller Lists
- 5. Accessing Ocr F29june 2012 Question Paper Free and Paid eBooks
 - Ocr F29june 2012 Question Paper Public Domain eBooks
 - Ocr F29june 2012 Question Paper eBook Subscription Services
 - Ocr F29june 2012 Question Paper Budget-Friendly Options
- 6. Navigating Ocr F29june 2012 Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr F29june 2012 Question Paper Compatibility with Devices
 - Ocr F29june 2012 Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr F29june 2012 Question Paper
 - Highlighting and Note-Taking Ocr F29june 2012 Question Paper
 - Interactive Elements Ocr F29june 2012 Question Paper
- 8. Staying Engaged with Ocr F29june 2012 Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr F29june 2012 Question Paper
- 9. Balancing eBooks and Physical Books Ocr F29june 2012 Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr F29june 2012 Question Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr F29june 2012 Question Paper
 - Setting Reading Goals Ocr F29june 2012 Question Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr F29june 2012 Question Paper
 - Fact-Checking eBook Content of Ocr F29june 2012 Question Paper
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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