

- 99) Ionization energies allow chemists to determine chemical structures. The first two ionization energies of rubidium and strontium are given in the table.

| element   | first ionization energy (kJ/mol) | second ionization energy (kJ/mol) |
|-----------|----------------------------------|-----------------------------------|
| rubidium  | 403                              | 2633                              |
| strontium | 550                              | 1064                              |

- 100) Write an equation to represent the **second** ionization energy of strontium.

Include state symbols.

[2]

- 101) Why is the **first** ionization energy of strontium larger than the **first** ionization energy of rubidium?

Explain answer (as above) use appropriate chemical terms **specifically**.

[2]

- 102) Why is the **second** ionization energy of rubidium larger than the **second** ionization energy of strontium?

[2]

(Total 12)

# Ocr Chemistry 2014 F321 Question Paper

**JS Bruner**



**Ocr Chemistry 2014 F321 Question Paper:**

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