

Physics

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(Chapter 11) (Thermal Properties of Matter) (Exercises)

(Class 11)

Question 11.15:

Given below are observations on molar specific heats at room temperature of some common gases.

Gas	Molar specific heat (C_v) ($\text{cal mol}^{-1} \text{K}^{-1}$)
Hydrogen	4.87
Nitrogen	4.97
Oxygen	5.02
Nitric oxide	4.99
Carbon monoxide	5.01
Chlorine	6.17

The measured molar specific heats of these gases are markedly different from those for monatomic gases. Typically, molar specific heat of a monatomic gas is 2.92 cal/mol K . Explain this difference. What can you infer from the somewhat larger (than the rest) value for chlorine?

Answer 11.15:

The gases listed in the given table are diatomic. Besides the translational degree of freedom, they have other degrees of freedom (modes of motion).

Heat must be supplied to increase the temperature of these gases. This increases the average energy of all the modes of motion. Hence, the molar specific heat of diatomic gases is more than that of monatomic gases.

If only rotational mode of motion is considered, then the molar specific heat of a diatomic gas

$$= \frac{5}{2} R \Rightarrow \frac{5}{2} \times 1.98 = 4.95 \text{ cal mol}^{-1} \text{K}^{-1}$$

With the exception of chlorine, all the observations in the given table agree with $\left(\frac{5}{2}R\right)$.

This is because at room temperature, chlorine also has vibrational modes of motion besides rotational and translational modes of motion.

Question 11.16:

A child running a temperature of 101°F is given an antipyrin (i.e. a medicine that lowers fever) which causes an increase in the rate of evaporation of sweat from his body. If the fever is brought down to 98°F in 20 minutes, what is the average rate of extra evaporation caused, by the drug. Assume the evaporation mechanism to be the only way by which heat is lost. The mass of the child is 30 kg. The specific heat of human body is approximately the same as that of water, and latent heat of evaporation of water at that temperature is about 580 cal g^{-1} .

Answer 11.16:

Initial temperature of the body of the child, $T_1 = 101^\circ\text{F}$

Final temperature of the body of the child, $T_2 = 98^\circ\text{F}$

$$\text{Change in temperature, } \Delta T = \left[(101 - 98) \times \frac{5}{9} \right]^\circ\text{C}$$

Time taken to reduce the temperature, $t = 20 \text{ min}$

Mass of the child, $m = 30 \text{ kg} = 30 \times 10^3 \text{ g}$

Specific heat of the human body = Specific heat of water = $c = 1000 \text{ cal/kg/}^\circ\text{C}$

Latent heat of evaporation of water, $L = 580 \text{ cal g}^{-1}$

The heat lost by the child is given as: $\Delta\theta = mc\Delta T = 30 \times 1000 \times (101 - 98) \times \frac{5}{9} = 50000 \text{ cal}$

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Table of Contents Ncert 11th Physics Notes Full Chapters

1. Understanding the eBook Ncert 11th Physics Notes Full Chapters
 - The Rise of Digital Reading Ncert 11th Physics Notes Full Chapters
 - Advantages of eBooks Over Traditional Books
2. Identifying Ncert 11th Physics Notes Full Chapters
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ncert 11th Physics Notes Full Chapters
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ncert 11th Physics Notes Full Chapters
 - Personalized Recommendations
 - Ncert 11th Physics Notes Full Chapters User Reviews and Ratings
 - Ncert 11th Physics Notes Full Chapters and Bestseller Lists

5. Accessing Ncert 11th Physics Notes Full Chapters Free and Paid eBooks
 - Ncert 11th Physics Notes Full Chapters Public Domain eBooks
 - Ncert 11th Physics Notes Full Chapters eBook Subscription Services
 - Ncert 11th Physics Notes Full Chapters Budget-Friendly Options
6. Navigating Ncert 11th Physics Notes Full Chapters eBook Formats
 - ePub, PDF, MOBI, and More
 - Ncert 11th Physics Notes Full Chapters Compatibility with Devices
 - Ncert 11th Physics Notes Full Chapters Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ncert 11th Physics Notes Full Chapters
 - Highlighting and Note-Taking Ncert 11th Physics Notes Full Chapters
 - Interactive Elements Ncert 11th Physics Notes Full Chapters
8. Staying Engaged with Ncert 11th Physics Notes Full Chapters
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ncert 11th Physics Notes Full Chapters
9. Balancing eBooks and Physical Books Ncert 11th Physics Notes Full Chapters
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ncert 11th Physics Notes Full Chapters
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ncert 11th Physics Notes Full Chapters
 - Setting Reading Goals Ncert 11th Physics Notes Full Chapters
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ncert 11th Physics Notes Full Chapters
 - Fact-Checking eBook Content of Ncert 11th Physics Notes Full Chapters
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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