

Worksheet #3: Naming Inorganic Compounds

1. Give the name for each of the following ionic compounds.

- a. AlF_3 Aluminum Fluoride
- b. Fe(OH)_2 Iron(II) Hydroxide
- c. $\text{Cu(NO}_3)_2$ Copper (II) Nitrate
- d. $\text{Ba(ClO}_4)_2$ Barium Perchlorate
- e. LiPO_3 Lithium Phosphite
- f. Hg_2S Mercury (I) Sulfide
- g. $\text{Ca(CH}_3\text{COO)}_2$ Calcium Acetate
- h. Cr_2CO_3 Chromium (III) Carbonate
- i. $\text{K}_2\text{Cr}_2\text{O}_7$ Potassium Dichromate
- j. NH_4HSO_4 Ammonium Sulfate

② all of the underlined compounds are polyatomic ions $\rightarrow$ Memorise These

2. Write the chemical formula for each of the following compounds:

- a. copper (II) oxide Cu_2O
 b. potassium peroxide K_2O_2
 c. aluminum hydroxide $\text{Al}(\text{OH})_3$
 d. iron(II) carbonate FeCO_3

Figure 13. 2010-2011 Season: Catches of 1

- a. mercury(I) bromide Hg_2Br_2
- b. iron(III) carbonate $\text{Fe}_2(\text{CO}_3)_3$
- c. sodium hypobromite NaBrO not
- most common polyatomic ions
 BrO^{1-}

Ground concentration: polyphosphate: 1 ppm
 $\text{B} = \text{O}^{2-}$ 10

2. Give the name or chemical formula, as appropriate, for each of the following acids:

- HNO_3 Nitric Acid
- HNO_2 Hydro Nitric Acid
- H_3PO_4 Phosphoric Acid
 phosphate^{3-} (the = oxy)

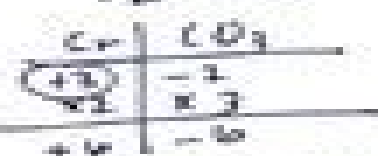
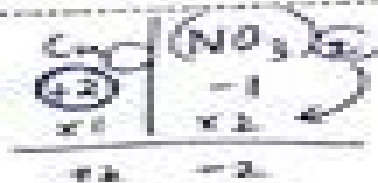
- d. hypochlorous acid HClO
e. iodic acid HIO_3
f. sulfurous acid H_2SO_3
g. break off of iodate as IO_3^-

^a Based off of Volume up TO₁

4. Give the name or chemical formula, as appropriate, for each of the following molecular substances:

1. Iodine hexafluoride
2. Iodine pentafluoride
3. Xenon Trioxide

- d. dinitrogen tetroxide: N_2O_4
e. hydrogen cyanide: HCN
f. tetraphosphorus hexasulfide: P_4S_6



— *W. J. G. S. J.*



1. The first step is to identify the problem.
 2. The second step is to define the problem.
 3. The third step is to analyze the problem.
 4. The fourth step is to develop a solution.
 5. The fifth step is to implement the solution.
 6. The sixth step is to evaluate the solution.
 7. The seventh step is to monitor the solution.
 8. The eighth step is to maintain the solution.
 9. The ninth step is to improve the solution.
 10. The tenth step is to document the solution.



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