Evaluate terms

(Volume of treatment tooks) = 
$$\left(\frac{1}{3}day\right)\left(\frac{1}{22000}\frac{m^3}{day}\right) = 10667 m^3$$

Thus the rate of reaction

So 
$$= \frac{dN_{e_{20}H_{NL}}}{dt} = \left(\frac{5400000 \, \text{ks/say}}{0.282 \, \text{kg/day}}\right) \left(\frac{da_{\gamma}}{24(3400)c}\right) (0.6) = 193 \, \text{well readed/c}$$

Thus the rate of disappearance of CaoHua

$$-V' = \frac{1}{W_{\text{res}}} \frac{dN}{dt} = \frac{1}{50000} \left( 135 \right) = 0.0027 \text{ mol/kg cat} \cdot 5$$

$$-V'' = \frac{1}{V_{\text{res}}} \frac{dN}{dt} = \frac{1}{625} \left( 125 \right) = 2.19 \text{ mol/m}^3 \text{cat} \cdot 5$$

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