

Chemistry Chapter 8 Test

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- ____ 1. What type of bond would you expect in a compound of oxygen and sodium?
 (a) ionic c. covalent
 b. metallic d. polar covalent
- ____ 2. The symbols δ^- and δ^+ indicate which of the following?
 a. electronegativity c. charge
 b. electrostatic force (d) partial charge
- ____ 3. Which of the following represents a nonpolar covalent molecule?
 a. NaCl c. HCN
 (b) Br₂ d. Au
- ____ 4. The measure of an atom's ability to attract electrons to itself in a chemical bond is called _____.
 a. electron affinity (c) electronegativity
 b. polarity d. ionization energy
- ____ 5. The polarity of a bond depends on the difference in the atoms' _____.
 a. electron affinities c. atomic radii
 (b) electronegativities d. ionization energies
- ____ 6. What type of bond would you expect to form between nitrogen and potassium?
 (a) ionic c. polar covalent
 b. covalent d. metallic
- ____ * High electronegativity differences are found in _____ bonds.
 a. metallic c. covalent
 (b) ionic metal - nonmetal d. Pauling
- ____ * 8. To achieve an octet when combining with potassium, would an atom of a Group 16 element be expected to lose, gain, or share electrons?
 a. lose K
 (b) gain K 17
 c. share
 (d) This cannot be determined from the given information. x
- ____ 9. Which of the following does not exist as a diatomic molecule?
 (a) argon c. chlorine
 b. hydrogen d. oxygen
 H N O F Cl Br I
 Gas: H₂ N₂ O₂ F₂ Cl₂ Br₂ I₂
- ____ 10. Which of these is not a covalent molecule?
 a. F₂ c. LiF
 b. CO₂ d. NH₃
- ____ 11. A dash in a Lewis structure represents _____.
 a. one electron c. four electrons
 (b) two electrons d. unpaired electrons
- ____ 12. The central atom in a Lewis structure is usually the atom _____.

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