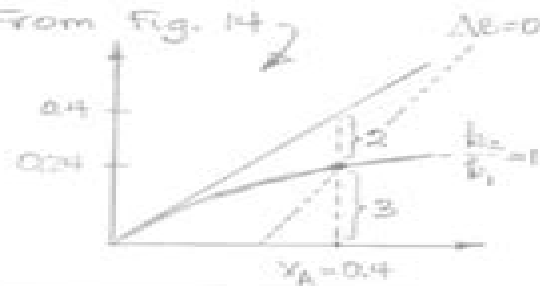


8.9

From Fig. 14



Thus

$$\text{given} \rightarrow \left. \begin{aligned} \frac{C_A}{C_{A0}} &= 0.6 \\ \frac{C_B}{C_{B0}} &= 0.24 \\ \frac{C_C}{C_{C0}} &= 0.16 \\ \frac{C_D}{C_{D0}} &= 0.42 \end{aligned} \right\} \frac{k_2}{k_1} = 1 \leftarrow$$

8.11

Evaluate k_1 and k_2 at various pH.

pH	k_1	k_2
2	3	11
4	5	7
6	7	11

From the table:

Operate at pH = 6 because
it gives the highest k_1 and the highest k_2 $\leftarrow$

8.13

Assume that $C_{A0} = 100$. Then

a) If you have produced lots of S & U, and if the second step reactions are relatively very slow, then the first step reactions are complete. So

$$C_A = 0 \quad C_R = 20 \quad C_S = 20 \quad C_T = 20 \quad \text{and} \quad C_U = 40 \leftarrow a)$$

b) The second step reactions are very fast, so any R formed has transformed into S and U. So $C_{R \text{ formed}} = 20 + 40 = 60$

Thus $C_T = \frac{1}{2} C_{R \text{ formed}} = 30$. So

$$C_A = 10 \quad C_R = 0 \quad C_S = 20 \quad C_T = 30 \quad \text{and} \quad C_U = 40 \leftarrow b)$$

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