

| Question |   |       | Answer                                                                                                                                                                                        | M                    | Guidance                                                                                                                                                                                                                             |
|----------|---|-------|-----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| 6        | a | i     | $\lambda$ is the (de Broglie) wavelength associated with a particle<br>$h$ is (a constant known as) the Planck constant<br>$mv$ is the momentum/mass $\times$ speed of the particle           | B1<br>B1<br>B1       | allow object/electron for particle<br>accept velocity for speed<br>accept $m$ and $v$ separately but the particle must appear with one of the quantities to score the mark                                                           |
|          |   | ii    | electrons passing through a (thin) sheet of graphite are diffracted producing rings/pattern (on a fluorescent screen)                                                                         | B1<br>A1             | any suitable situation <b>NOT</b> Au or Al, etc.<br>what is observed + interpretation 1 mark.<br><b>Do not</b> award the A mark alone unless a plausible situation has been suggested.<br>Young slits type experiments etc score 0/2 |
|          |   | iii 1 | $eV = \frac{1}{2} mv^2$<br>$2eV/m = v^2$<br>so $v = \sqrt{(2 \times 1.6 \times 10^{-19} \times 5.0 \times 10^3 / 9.1 \times 10^{-31})}$<br>so $v = 1.3 \times 10^8 \text{ (m s}^{-1}\text{)}$ | C1<br>C1<br>B1<br>A0 | allow $eV = 8.0 \times 10^{-17} \text{ J}$ for 1 mark<br>allow $\sqrt{(2 \times 8.0 \times 10^{-17} / 9.1 \times 10^{-31})}$ or evidence of correct calculation, e.g. $v = 1.325 \times 10^8$                                        |
|          |   | iii 2 | $\lambda = h/mv = 6.63 \times 10^{-34} / 1.3 \times 10^8 \times 9.1 \times 10^{-31}$<br>$\lambda = 5.6 \times 10^{-12} \text{ (m)}$                                                           | C1<br>A1             | allow $5.5 \times 10^{-12}$ if $1.325 \times 10^8$ is used                                                                                                                                                                           |
|          |   | iii 3 | $\lambda$ of visible light $5.6 \times 10^{-7} \text{ m}$<br>so power of 10 = 5 or -5                                                                                                         | B1<br>B1             | ect all 2; accept 4 to $7 \times 10^{-7} \text{ m}$ <b>NOT</b> $10^{-7}$<br><b>NOT</b> $10^5$ on the answer line                                                                                                                     |
|          | b |       | The photoelectric effect                                                                                                                                                                      | B1                   | QWC mark allow exp't description: uv light shone on a zinc plate connected to a gle<br>max 3 from 6 marking points                                                                                                                   |
|          |   |       | individual photons are absorbed by individual electrons in the metal surface, i.e. one to one interaction.                                                                                    | B1                   |                                                                                                                                                                                                                                      |
|          |   |       | Only photons with energies above the work function energy will cause photoelectron emission/idea of threshold frequency                                                                       | B1                   |                                                                                                                                                                                                                                      |
|          |   |       | Hence u-v photons or blue photons will cause photoemission but red photons will not.                                                                                                          | B1                   |                                                                                                                                                                                                                                      |
|          |   |       | Number of electrons emitted depends on light intensity                                                                                                                                        | B1                   |                                                                                                                                                                                                                                      |
|          |   |       | A wave model does not explain instantaneous emission of electrons.                                                                                                                            | B1                   |                                                                                                                                                                                                                                      |
|          |   |       | A wave model does not explain a threshold frequency/wavelength for emission to occur                                                                                                          | B1                   | allow wave model does not explain no emission however bright the light if energy of photon below work function                                                                                                                       |
|          |   |       | <b>Total question 6</b>                                                                                                                                                                       | <b>16</b>            |                                                                                                                                                                                                                                      |

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