

| Topic                | Grade | What do you need to know                                                                                                                                                                                                                                                                                       |  or  | What do You have to work on? |
|----------------------|-------|----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------------------------------------------------------------------------|------------------------------|
| Atomic Structure     | A     | <ul style="list-style-type: none"> <li>Why an atom is neutral in terms of sub-atomic particles</li> <li>Outline isotopes from data about numbers of protons and neutrons</li> </ul>                                                                                                                            |                                                                                                                                                                        |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Know the charges and masses of protons, electrons and neutrons.</li> <li>Know how isotopes are the same but have different mass number.</li> </ul>                                                                                                                      |                                                                                                                                                                        |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Describe the arrangement of an atoms electron structure - 2.8.8.2</li> <li>Describe the arrangement of elements in the periodic table.</li> </ul>                                                                                                                       |                                                                                                                                                                        |                              |
| Ionic Bonding        | A     | <ul style="list-style-type: none"> <li>Explain ionic bonding using the 'dot and cross' model.</li> <li>Explain the melting point of NaCl by the structure of and bonding.</li> <li>Explain why the melting point of NaCl is lower than that of MgO.</li> </ul>                                                 |                                                                                                                                                                        |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Predict the properties of substances with giant ionic structures.</li> <li>Understand that atoms with a full outer shell are electronically stable.</li> </ul>                                                                                                          |                                                                                                                                                                        |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Ionic bonds are formed from metals and non-metals</li> <li>Electrons are donated and received.</li> <li>Metals always donate the electron(s).</li> <li>Non- metals always receive the electron(s)</li> </ul>                                                            |                                                                                                                                                                        |                              |
| The Periodic Table   | A     | <ul style="list-style-type: none"> <li>Explain some covalent bonds using a 'dot and cross' model</li> <li>Explain double covalent binds using a 'dot and cross' model</li> </ul>                                                                                                                               |                                                                                                                                                                        |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Know why CO<sub>2</sub> and H<sub>2</sub>O have low melting points</li> </ul>                                                                                                                                                                                           |                                                                                                                                                                        |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Covalent bonds always have non-metals and share electrons</li> <li>Know that the group number is the same as the number of electrons in the outer shell.</li> <li>Be able to find out how many shells an atom has from the period it is in.</li> </ul>                  |                                                                                                                                                                        |                              |
| The Group 1 Elements | A     | <ul style="list-style-type: none"> <li>Construct a balanced symbol equation for Cu<sup>2+</sup>, Fe<sup>2+</sup> and Fe<sup>3+</sup> to make Cu(OH)<sub>2</sub>, Fe(OH)<sub>2</sub>, Fe(OH)<sub>3</sub>.</li> <li>Reactivity with water gives NaOH and KOH.</li> <li>Know the process is oxidation.</li> </ul> |                                                                                                                                                                        |                              |
|                      | B     | <ul style="list-style-type: none"> <li>Why group 1 elements form similar positive ions.</li> <li>Construct a balance symbol equation of group 1 with water</li> </ul>                                                                                                                                          |                                                                                                                                                                        |                              |
|                      | C     | <ul style="list-style-type: none"> <li>Metal + Water <math>\longrightarrow</math> Metal Hydroxide + Hydrogen</li> <li>Know how to use a flame test to identify potassium, lithium and sodium compounds.</li> </ul>                                                                                             |                                                                                                                                                                        |                              |
| The Group 7 Elements | A     | <ul style="list-style-type: none"> <li>Predict if displacement reactions will occur with other halogens</li> <li>Explain the trend of reactivity of group 7 by ideas of electron gain</li> <li>Explain why a process is reduced from its ionic</li> </ul>                                                      |                                                                                                                                                                        |                              |

# Ocr Gateway Separate Sciences B C4 Checklist

**Sam Holyman**



## **Ocr Gateway Separate Sciences B C4 Checklist:**

**OCR Twenty First Century Science** Dorothy Warren,2006-06 A revision guide that covers the core content of the OCR Science A single award specification from the Twenty First Century Science Suite **Separate Sciences B** ,2011 **OCR Gateway Science** ,2006 **OCR gateway science** Chris Sherry,2005 *OCR Gateway GCSE Separate Sciences Student Book* Graham Bone,Simon Broadley,Sue Hocking,Mark Matthews,Jim Newall,Angela Saunders,Nigel Saunders,2011-06-09 Our OCR GCSE Gateway Science course provides the simplest clearest approach to the new specifications easing the transition for you and helping students understand how to maximise their potential We focus on support motivation and exam preparation The Student Book helps your students develop clear scientific knowledge and understanding *OCR Gateway GCSE Separate Sciences Resources and Planning Pack* Chadha,2011-07-21 Our OCR GCSE Gateway Science course provides the simplest clearest approach to the new specifications easing the transition for you and helping students understand how to maximise their potential We focus on support motivation and exam preparation This pack provides lesson plans homework activity sheets and teacher guidance **OCR Gateway Science B** Natalie King,Samantha Holyman,Claire Hutchinson,2011-09 With clear and concise revision notes that cover everything you ll need to know for the exam this effective OCR GCSE Science revision guide supports active revision for students working at all levels Plenty of quick tests and practice questions test and reinforce understanding of the key content Providing a student friendly uncluttered approach to GCSE revision this OCR Gateway Science B revision guide is suitable for all levels and contains succinct revision notes and practice questions that focus on the core content needed for the exams Included in this book clear and concise coverage of all the exam assessed content simple and engaging explanations quick tests and practice questions throughout to test and reinforce understanding key words and supporting glossary Separate units differentiated by colour and labelling exam style questions at the end of each unit overview of How Science Works ence Worksence Worksence Works *Gateway Science for OCR GCSE Biology, Chemistry, Physics B* ,2006 **Gateway Science for OCR GCSE Additional Science B** ,2006

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