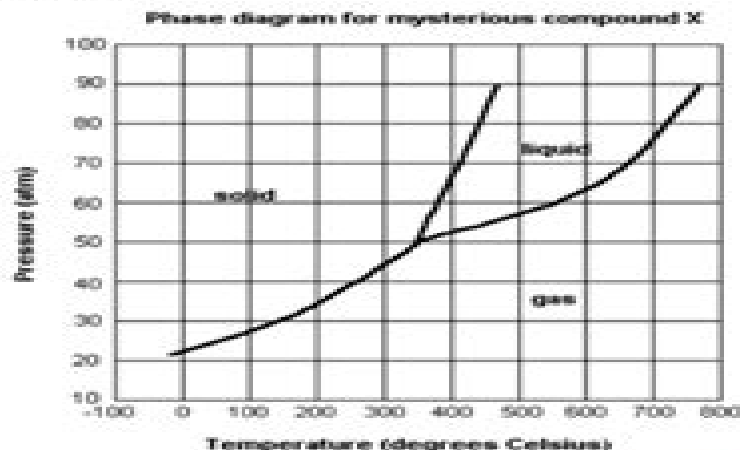


## Phase Diagram Worksheet

For each of the questions on this worksheet, refer to the phase diagram for mysterious compound X.



- 1) What is the critical temperature of compound X?  **$\sim 470^{\circ}\text{C}$**
- 2) If you were to have a bottle containing compound X in your closet, what phase would it most likely be in?  
**Extrapolating from this diagram, it's most likely a gas.**
- 3) At what temperature and pressure will all three phases coexist?  
 **$350^{\circ}\text{C}$ ,  $\sim 51\text{ atm}$**
- 4) If I have a bottle of compound X at a pressure of 45 atm and temperature of  $100^{\circ}\text{C}$ , what will happen if I raise the temperature to  $400^{\circ}\text{C}$ ?  
**It will sublime**
- 5) Why can't compound X be boiled at a temperature of  $200^{\circ}\text{C}$ ?  
**It does not form a liquid at this temperature. It only exists as a liquid at temperatures above  $350^{\circ}\text{C}$ .**
- 6) If I wanted to, could I drink compound X?  
**No. At the temperatures and pressures that it forms a liquid, you'd probably die.**

# Phase Diagram Practice Answer Key

**TD Snyder**



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