Name:	Date:	

Student Exploration: Phase Changes (4b)

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: altitude, boil, boiling point, freeze, freezing point, gas, liquid, melt, melting point, phase, solid

Activity B:	Get the Gizmo ready:	0-000
Temperature and molecular motion	 Click Reset, and select the Micro view. Set Ice volume to 0 cm³. Set Add/rem ove heat energy to 0 J/s. 	8 6 8 9

Question: Why do phase changes occur?

Compare: Set the Water temperature to 0 °C and click Play. Observe the water molecules. Click Reset, set the Water temperature to 100 °C, and click Play again.

What do you notice?

There is no difference.

- Observe: Click Reset. The mean molecular speed of the water molecules is displayed below the container. Set the Water temperature to 0 °C and Add/remove heat energy to 400 J/s. Click Play.
 - A. How does the mean speed of the water molecules change as they are heated?

The water gets hotter (has more heat energy).

B. Does the mean molecular speed change as much as the temperature as the water heats up? Explain.

Yes, the more the temperature increases the more the molecules of the water speed up (until they hit the boiling point, where they cannot get hotter any longer.)

3. Explain: How is temperature related to the motions of molecules?

The higher the temperature, the faster the molecules move.

4. Observe: Click Reset. Set the Water temperature to 20 °C and the Ice volume to 50 cm³. Set Add/rem ove heat energy to 0 J/s. Click Play. How do the molecules in the liquid interact with the molecules in the solid?

Phase Changes Answer Key

Daniela Niemeyer

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Table of Contents Phase Changes Answer Key

- 1. Understanding the eBook Phase Changes Answer Key
 - The Rise of Digital Reading Phase Changes Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Phase Changes Answer Key
 - Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Phase Changes Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Phase Changes Answer Key

- Personalized Recommendations
- Phase Changes Answer Key User Reviews and Ratings
- Phase Changes Answer Key and Bestseller Lists
- 5. Accessing Phase Changes Answer Key Free and Paid eBooks
 - Phase Changes Answer Key Public Domain eBooks
 - Phase Changes Answer Key eBook Subscription Services
 - Phase Changes Answer Key Budget-Friendly Options
- 6. Navigating Phase Changes Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Phase Changes Answer Key Compatibility with Devices
 - Phase Changes Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Phase Changes Answer Key
 - Highlighting and Note-Taking Phase Changes Answer Key
 - Interactive Elements Phase Changes Answer Key
- 8. Staying Engaged with Phase Changes Answer Key
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Phase Changes Answer Key
- 9. Balancing eBooks and Physical Books Phase Changes Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Phase Changes Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Phase Changes Answer Key
 - Setting Reading Goals Phase Changes Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Phase Changes Answer Key

- Fact-Checking eBook Content of Phase Changes Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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