

pH AND pOH

Name _____

The pH of a solution indicates how acidic or basic that solution is.

• pH range of 0 - 7 acidic

7 neutral

7-14 basic

Since $[H^+][OH^-] = 10^{-14}$ at $25^\circ C$, if $[H^+]$ is known, the $[OH^-]$ can be calculated and vice versa.

$$pH = -\log [H^+]$$

$$\text{So if } [H^+] = 10^{-6} M, pH = 6.$$

$$pOH = -\log [OH^-]$$

$$\text{So if } [OH^-] = 10^{-8} M, pOH = 8.$$

$$\text{Together, } pH + pOH = 14.$$

Complete the following chart.

	$[H^+]$	pH	$[OH^-]$	pOH	Acidic or Basic
1.	$10^{-5} M$	5	$10^{-9} M$	9	Acidic
2.		7			
3.			$10^{-4} M$		
4.	$10^{-2} M$				
5.				11	
6.		12			
7.			$10^{-5} M$		
8.	$10^{-11} M$				
9.				13	
10.		6			

Ph And Poh Answer Key

JG Myers



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Table of Contents Ph And Poh Answer Key

1. Understanding the eBook Ph And Poh Answer Key
 - The Rise of Digital Reading Ph And Poh Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Ph And Poh Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ph And Poh Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ph And Poh Answer Key
 - Personalized Recommendations
 - Ph And Poh Answer Key User Reviews and Ratings
 - Ph And Poh Answer Key and Bestseller Lists
5. Accessing Ph And Poh Answer Key Free and Paid eBooks
 - Ph And Poh Answer Key Public Domain eBooks
 - Ph And Poh Answer Key eBook Subscription Services

- Ph And Poh Answer Key Budget-Friendly Options
- 6. Navigating Ph And Poh Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Ph And Poh Answer Key Compatibility with Devices
 - Ph And Poh Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ph And Poh Answer Key
 - Highlighting and Note-Taking Ph And Poh Answer Key
 - Interactive Elements Ph And Poh Answer Key
- 8. Staying Engaged with Ph And Poh Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ph And Poh Answer Key
- 9. Balancing eBooks and Physical Books Ph And Poh Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ph And Poh Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ph And Poh Answer Key
 - Setting Reading Goals Ph And Poh Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ph And Poh Answer Key
 - Fact-Checking eBook Content of Ph And Poh Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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