

## **Calculating pH of Acids – Practice**

**Instructions:** Use the equations for pH to complete the problems below.

- 1) What is the pH of a 0.0266 M hydrobromic acid (HBr) solution?

$$\text{pH} = -\log[\text{H}^+] \rightarrow \text{pH} = -\log[0.0266] \rightarrow \boxed{\text{pH} = 1.58}$$

- 2) Calculate the hydrogen ion concentration of a solution with a pH of 6.3.

$$[\text{H}^+] = 10^{-\text{pH}} \rightarrow [\text{H}^+] = 10^{-6.3} \rightarrow \boxed{[\text{H}^+] = 5.01 \times 10^{-7}}$$

- 3) Calculate the pH of a 0.004 M hydrofluoric acid (HF) solution?

$$\text{pH} = -\log[\text{H}^+] \rightarrow \text{pH} = -\log[0.004] \rightarrow \boxed{\text{pH} = 2.40}$$

- 4) What is the hydrogen ion concentration of an acid with a pH of 2.21?

$$[\text{H}^+] = 10^{-\text{pH}} \rightarrow [\text{H}^+] = 10^{-2.21} \rightarrow \boxed{[\text{H}^+] = 6.17 \times 10^{-3}}$$

- 5) Calculate the pH of a 1.14 L solution that contains 0.19 moles of HCl.

$$[\text{H}^+] = \frac{\text{moles of solute}}{\text{liters of solvent}} \rightarrow \frac{0.19 \text{ mol}}{1.14 \text{ L}} \rightarrow 0.167 \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] \rightarrow \text{pH} = -\log[0.167] \rightarrow \boxed{\text{pH} = 0.78}$$

# Ph Review Problems Worksheet Answers

**David Baud**



## **Ph Review Problems Worksheet Answers:**

## **Ph Review Problems Worksheet Answers** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Ph Review Problems Worksheet Answers**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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### **Table of Contents Ph Review Problems Worksheet Answers**

1. Understanding the eBook Ph Review Problems Worksheet Answers
  - The Rise of Digital Reading Ph Review Problems Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ph Review Problems Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in a Ph Review Problems Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ph Review Problems Worksheet Answers
  - Personalized Recommendations
  - Ph Review Problems Worksheet Answers User Reviews and Ratings

- Ph Review Problems Worksheet Answers and Bestseller Lists
- 5. Accessing Ph Review Problems Worksheet Answers Free and Paid eBooks
  - Ph Review Problems Worksheet Answers Public Domain eBooks
  - Ph Review Problems Worksheet Answers eBook Subscription Services
  - Ph Review Problems Worksheet Answers Budget-Friendly Options
- 6. Navigating Ph Review Problems Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Ph Review Problems Worksheet Answers Compatibility with Devices
  - Ph Review Problems Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ph Review Problems Worksheet Answers
  - Highlighting and Note-Taking Ph Review Problems Worksheet Answers
  - Interactive Elements Ph Review Problems Worksheet Answers
- 8. Staying Engaged with Ph Review Problems Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ph Review Problems Worksheet Answers
- 9. Balancing eBooks and Physical Books Ph Review Problems Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ph Review Problems Worksheet Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ph Review Problems Worksheet Answers
  - Setting Reading Goals Ph Review Problems Worksheet Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ph Review Problems Worksheet Answers
  - Fact-Checking eBook Content of Ph Review Problems Worksheet Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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