

7. A student's automobile uses 1/3 gallon (1160 grams) of octane on each round trip to and from school. What volume of CO_2 gas is produced at -15°C and 29.55 in Hg?

$$2\text{C}_8\text{H}_{18} + 25\text{O}_2 \rightarrow 16\text{CO}_2 + 18\text{H}_2\text{O}$$

$$\frac{1160\text{g}}{114\text{g/mol}} = 10.2\text{mol} \cdot \frac{16\text{mol CO}_2}{2\text{mol}} = 81.4\text{mol CO}_2 \quad V = \frac{nRT}{P} = 1670\text{ L CO}_2$$

8. Given: $2\text{NH}_4\text{Cl}_{(\text{aq})} + \text{Ca}(\text{OH})_{2(\text{aq})} \rightarrow \text{CaCl}_{2(\text{aq})} + 2\text{NH}_{3(\text{g})} + 2\text{H}_2\text{O}_{(\text{l})}$. What volume of ammonia gas is produced at 72.0°C and 3.55 atm when 6.20g of ammonium chloride reacts with 9.00g of calcium hydroxide?

$$\begin{array}{l} \frac{6.20\text{g}}{53.5\text{g/mol}} = 0.116\text{mol} \\ \frac{9.00\text{g}}{74\text{g/mol}} = 0.122\text{mol} \end{array}$$

$$0.116\text{mol} \xrightarrow{\text{L.R.}} 0.058\text{mol} \cdot \frac{2\text{mol NH}_3}{1\text{mol}} = 0.116\text{mol NH}_3$$

$$V = \frac{nRT}{P} = \frac{0.116\text{mol} \cdot 0.0821\text{ L}\cdot\text{atm/mol}\cdot\text{K} \cdot 344\text{K}}{3.55\text{ atm}} = 0.863\text{ L NH}_3$$

9. From the decomposition of 2.00grams of CaCO_3 a student collects 358 ml of CO_2 gas over water at 19.0°C and at a pressure of 752 torr. What is the percent yield? V.P._{19.0} at 19.0°C = 16.6 mm Hg

$$\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$$

$$\frac{2.00\text{g}}{100} = 0.02\text{mol CaCO}_3 \rightarrow 0.02\text{mol CO}_2$$

$$P = 0.752\text{ atm} - 0.0216\text{ atm} = 0.730\text{ atm}$$

$$V = 0.358\text{ L}$$

$$n = \frac{PV}{RT} = 0.0145\text{mol CO}_2$$

$$T = 292\text{ K}$$

$$\frac{0.636\text{g}}{0.870\text{g}} \cdot 100 = 72.3\%$$

10. 260.ml of a solution of nitric acid, HNO_3 , is analyzed with an excess of CaCO_3 by the reaction:



If 72.0 ml of CO_2 gas is produced at 23.0°C and 29.38 in Hg, what is the concentration of the original nitric acid in units of mg HNO_3 /ml of solution?

$$n = \frac{P \cdot V}{RT} = 2.91 \times 10^{-3}\text{mol CO}_2$$

$$2.91 \times 10^{-3}\text{mol CO}_2 \cdot \frac{2\text{mol H}^+}{1\text{mol CO}_2} \cdot \frac{1\text{mol HNO}_3}{1\text{mol H}^+} \cdot \frac{63\text{g}}{1\text{mol HNO}_3} \cdot \frac{1000\text{mg}}{1\text{g}} = \frac{366.7\text{mg HNO}_3}{260.\text{mL}} = 1.41\text{ mg/mL}$$

11. Consider the following unbalanced equation: $\text{Cu}_2\text{S}_{(\text{s})} + \text{O}_{2(\text{g})} \rightarrow \text{Cu}_2\text{O}_{(\text{s})} + \text{SO}_{2(\text{g})}$. If 1.22 grams of Cu_2S reacted with 4.52 liters of O_2 gas at 28°C and 1.0 atm what volume of SO_2 gas would be formed at $100.^\circ\text{C}$ and 0.933 atm?

$$2\text{Cu}_2\text{S} + 3\text{O}_2 \rightarrow 2\text{Cu}_2\text{O} + 2\text{SO}_2$$

$$\frac{1.22\text{g}}{157.6\text{g/mol}} = 7.66 \times 10^{-3}\text{mol}$$

$$n = \frac{PV}{RT} = 0.183\text{mol}$$

$$V = \frac{nRT}{P} = 0.252\text{ L SO}_2$$

$$7.66 \times 10^{-3}\text{mol Cu}_2\text{S} \cdot \frac{2\text{mol SO}_2}{2\text{mol Cu}_2\text{S}} = 7.66 \times 10^{-3}\text{mol SO}_2$$

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**Thirumalaisamy P.
Velavancorresponding**



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