

ACTIVITY: Molecular Polarity Computer Simulation (PhET)

Go to: <https://phet.colorado.edu/> Click on the arrow, DOWNLOAD, OPEN and RUN the MOLECULE POLARITY program. Or google "phet molecule polarity" to find the simulation.

Introduction:

In this atomic level simulation, you will investigate how atoms' **electronegativity** values affect the bonds they produce. When two atoms form a covalent bond, a pair of electrons is shared between atoms. **Electronegativity** is a measure of a single atom's ability to attract electrons shared in that bond. In this lab you will work at your own pace, to answer a number of questions. To begin, from what you've already learned about the protons and electrons in an atom, what would cause an atom to have a **high** electronegativity value?

Why might an atom have a low electronegativity value?

Procedure:

Two Atoms

1. Turn on (check) all view options. **distill** out just the molecule with your mouse.
2. Take your time and investigate how the binary compound's bond behaves when the atom's electronegativity and orientation are changed. **Do not rush through this step.** Describe the bond formed between two atoms with **similar, low** electronegativities.

Describe the bond formed between two atoms with **similar, high** electronegativities.

Describe the bond formed between two atoms with **very different** electronegativities.

Describe (in your own words) what is meant by **partial charges**, δ^- and δ^+ .

δ^- represents: _____

δ^+ represents: _____

What happens when the electric field is applied to a very polar molecule?

Why do you think this is? _____

View

- ☒ Bond Dipoles **ON**
- ☒ Partial Charges
- ☒ Bond Orientation

Surfaces

- ☐ None
- ☐ Electrostatic Potential
- ☒ Electron Density

Electric Field

- ☒ ON ☐ OFF



Phet Molecular Polarity Lab Answer Key

Kathleen Armour



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