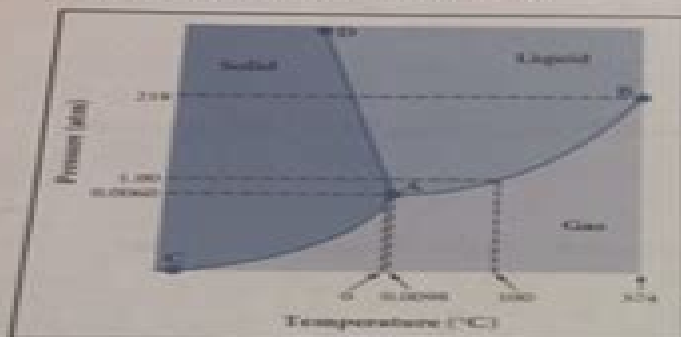
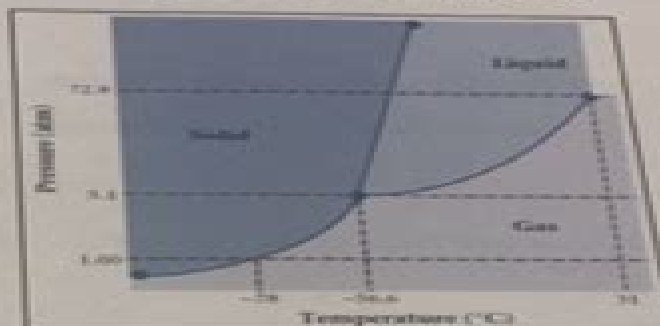


Investigating Phase Diagrams:

Consider the phase diagrams of water and carbon dioxide (Kelter: fig 11.18 and 11.19)

Phase Diagram for Water ( $H_2O$ )Phase Diagram for Carbon dioxide ( $CO_2$ )

1. What are the axis labels for a phase diagram? \_\_\_\_\_ versus \_\_\_\_\_.
2. Study some of the pressures and temperatures marked on the two diagrams. Are the diagrams drawn to scale (i.e. Are the axes linear)? \_\_\_\_\_

Consider the diagram for water:

3. What phase change occurs as the temperature changes from 100°C to 370°C at a constant applied pressure of 5 atm?

Information: The curve from A  $\rightarrow$  B represents vaporization (i.e. boiling)/condensation. This occurs when liquid and vapor are in equilibrium.

4. According to the phase diagram, can water boil at 50°C? \_\_\_\_\_. Does this occur above or below 1.00 atm?
5. Is it possible to melt ice at -10°C by increasing the pressure applied to the ice? \_\_\_\_\_

Consider the diagram for carbon dioxide:

6. Consider the  $CO_2$  phase diagram at a constant pressure of 1.00 atm and the temperature changes from -85°C to 0°C.
  - a) What are the initial and final  $CO_2$  phases present at -85°C \_\_\_\_\_ and 0°C \_\_\_\_\_
  - b) What term is used to describe this phase change? \_\_\_\_\_

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**KJ Lindholm-Leary**



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