

Calculating pH of Acids – Practice

Instructions: Use the equations for pH to complete the problems below.

- 1) What is the pH of a 0.0266 M hydrobromic acid (HBr) solution?

$$\text{pH} = -\log[\text{H}^+] \rightarrow \text{pH} = -\log[0.0266] \rightarrow \boxed{\text{pH} = 1.58}$$

- 2) Calculate the hydrogen ion concentration of a solution with a pH of 6.3.

$$[\text{H}^+] = 10^{-\text{pH}} \rightarrow [\text{H}^+] = 10^{-6.3} \rightarrow \boxed{[\text{H}^+] = 5.01 \times 10^{-7}}$$

- 3) Calculate the pH of a 0.004 M hydrofluoric acid (HF) solution?

$$\text{pH} = -\log[\text{H}^+] \rightarrow \text{pH} = -\log[0.004] \rightarrow \boxed{\text{pH} = 2.40}$$

- 4) What is the hydrogen ion concentration of an acid with a pH of 2.21?

$$[\text{H}^+] = 10^{-\text{pH}} \rightarrow [\text{H}^+] = 10^{-2.21} \rightarrow \boxed{[\text{H}^+] = 6.17 \times 10^{-3}}$$

- 5) Calculate the pH of a 1.14 L solution that contains 0.19 moles of HCl.

$$[\text{H}^+] = \frac{\text{moles of solute}}{\text{liters of solvent}} \rightarrow \frac{0.19 \text{ mol}}{1.14 \text{ L}} \rightarrow 0.167 \text{ M}$$

$$\text{pH} = -\log[\text{H}^+] \rightarrow \text{pH} = -\log[0.167] \rightarrow \boxed{\text{pH} = 0.78}$$

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P. H. Fung



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Table of Contents Ph Calculations Worksheet Chemfiesta Answers

1. Understanding the eBook Ph Calculations Worksheet Chemfiesta Answers
 - The Rise of Digital Reading Ph Calculations Worksheet Chemfiesta Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Ph Calculations Worksheet Chemfiesta Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ph Calculations Worksheet Chemfiesta Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ph Calculations Worksheet Chemfiesta Answers
 - Personalized Recommendations
 - Ph Calculations Worksheet Chemfiesta Answers User Reviews and Ratings
 - Ph Calculations Worksheet Chemfiesta Answers and Bestseller Lists
5. Accessing Ph Calculations Worksheet Chemfiesta Answers Free and Paid eBooks
 - Ph Calculations Worksheet Chemfiesta Answers Public Domain eBooks
 - Ph Calculations Worksheet Chemfiesta Answers eBook Subscription Services
 - Ph Calculations Worksheet Chemfiesta Answers Budget-Friendly Options
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- ePub, PDF, MOBI, and More
 - Ph Calculations Worksheet Chemfiesta Answers Compatibility with Devices
 - Ph Calculations Worksheet Chemfiesta Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ph Calculations Worksheet Chemfiesta Answers
 - Highlighting and Note-Taking Ph Calculations Worksheet Chemfiesta Answers
 - Interactive Elements Ph Calculations Worksheet Chemfiesta Answers
 8. Staying Engaged with Ph Calculations Worksheet Chemfiesta Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ph Calculations Worksheet Chemfiesta Answers
 9. Balancing eBooks and Physical Books Ph Calculations Worksheet Chemfiesta Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ph Calculations Worksheet Chemfiesta Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ph Calculations Worksheet Chemfiesta Answers
 - Setting Reading Goals Ph Calculations Worksheet Chemfiesta Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ph Calculations Worksheet Chemfiesta Answers
 - Fact-Checking eBook Content of Ph Calculations Worksheet Chemfiesta Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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