

The letters "pH" represent the French words "pouvoir hydrogene" which means "hydrogen power".

- The definition of pH is *pH is equal to the negative log (logarithm) of the hydrogen ion concentration of a solution.*
- The logarithm of a number is the power to which 10 must be raised to equal that number.

A pH value of less than 7 indicates a(n) ACIDIC solution. A pH value of 7 indicates a neutral solution. A pH value of more than 7 indicates a(n) BASIC solution.

**PROBLEMS:** Show all work and circle the final answer.

1. Determine the pH of a 0.010 M  $\text{HNO}_3$  solution.

$$\begin{aligned} \text{pH} &= -\log[0.010] \\ &= \boxed{2} \end{aligned}$$

2. What is the pH of a  $2.5 \times 10^{-6}$  M solution of HCl?

$$\begin{aligned} \text{pH} &= -\log[2.5 \times 10^{-6}] \\ &= \boxed{5.6} \end{aligned}$$

3. Calculate the pH of a solution of 0.0025M  $\text{H}_2\text{SO}_4$ .

$$\begin{aligned} \text{pH} &= -\log[2 \times 0.0025] \\ &= -\log[0.005] \\ &= \boxed{2.3} \end{aligned}$$

4. Calculate the pH of a 0.0010 M NaOH solution.

$$\begin{aligned} [\text{OH}^-][0.0010] &= 1.0 \times 10^{-14} \\ [\text{OH}^-] &= \frac{1.0 \times 10^{-14}}{0.0010} \end{aligned}$$

$$\text{pH} = -\log[1 \times 10^{-11}]$$

$$\text{pH} = \boxed{11}$$

5. What is the pH of a 0.020M  $\text{Sr}(\text{OH})_2$  solution?

$$[\text{OH}^-][0.04] = 1 \times 10^{-14}$$

$$\begin{aligned} \text{pH} &= -\log[\text{H}^+] \\ &= -\log 2.5 \times 10^{-9} \end{aligned}$$

# Ph Calculations Answers With Work Chemfiesta

**James Newton Butler**



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## **Table of Contents Ph Calculations Answers With Work Chemfiesta**

1. Understanding the eBook Ph Calculations Answers With Work Chemfiesta
  - The Rise of Digital Reading Ph Calculations Answers With Work Chemfiesta
  - Advantages of eBooks Over Traditional Books
2. Identifying Ph Calculations Answers With Work Chemfiesta
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ph Calculations Answers With Work Chemfiesta
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ph Calculations Answers With Work Chemfiesta
  - Personalized Recommendations
  - Ph Calculations Answers With Work Chemfiesta User Reviews and Ratings
  - Ph Calculations Answers With Work Chemfiesta and Bestseller Lists
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    - ePub, PDF, MOBI, and More
    - Ph Calculations Answers With Work Chemfiesta Compatibility with Devices
    - Ph Calculations Answers With Work Chemfiesta Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Ph Calculations Answers With Work Chemfiesta
    - Highlighting and Note-Taking Ph Calculations Answers With Work Chemfiesta
    - Interactive Elements Ph Calculations Answers With Work Chemfiesta
  8. Staying Engaged with Ph Calculations Answers With Work Chemfiesta
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Ph Calculations Answers With Work Chemfiesta
  9. Balancing eBooks and Physical Books Ph Calculations Answers With Work Chemfiesta
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Ph Calculations Answers With Work Chemfiesta
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Ph Calculations Answers With Work Chemfiesta
    - Setting Reading Goals Ph Calculations Answers With Work Chemfiesta
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Ph Calculations Answers With Work Chemfiesta
    - Fact-Checking eBook Content of Ph Calculations Answers With Work Chemfiesta
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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