# pH AND pOH

Name \_

The pH of a solution indicates how acidic or basic that solution is.

pH range of 0 - 7 acidic

7 neutral

7-14 basic

Since [H\*] [OH-] =  $10^{-14}$  at 25° C, if [H\*] is known, the [OH-] can be calculated and vice versa.

$$pH = -log \{H^*\}$$

$$pOH = -log[OH]$$

$$pOH = -log [OH-]$$
 So if  $[OH-] = 10^{+} M, pOH = 8.$ 

Together, pH + pOH = 14.

Complete the following chart.

	[H*]	pH	[OH-]	рОН	Acidic or Basic
1.	10° M	5	10° M	9	Acidic
2.		7			
3.			104 M		
4.	10° M				
5.				11	
6.		12			
7.			10° M		
8.	10-11 M				
9.				13	
10.		6			

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