

- (c) 350 cm, convergent
(d) 150 cm, divergent
16. Three resistances A, B and C each of 5Ω are connected to a battery of 10 V as shown in the figure. Calculate the current through C:
- 
- (a) 2 amp (b) 1 amp
(c) 3 amp (d) 0.6 amp
17. The device used to generate electrical energy is:
- (a) Electric motor (b) Generator
(c) Galvanometer (d) Voltmeter
18. The number of molecules in 11 g of CO_2 are:
- (a) 0.25×10^{23} (b) 0.50×10^{23}
(c) 1.00×10^{23} (d) 1.51×10^{23}
19. The value of charge / mass ratio of electron was determined by:
- (a) W.K. Roentgen (b) J.J. Thomson
(c) Marie Curie (d) Niels Bohr
20. Which of the following reactions is an example of combination reaction?
- (a) $\text{Zn(s)} + \text{CuSO}_4(\text{aq}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Cu(s)}$
(b) $\text{BaCl}_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq}) \rightarrow \text{BaSO}_4(\text{s}) + 2\text{NaCl}(\text{aq})$
(c) $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO(s)} + \text{CO}_2(\text{g})$
(d) $\text{C(s)} + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$
21. The element 'X' and 'Y' have atomic numbers 12 and 17 respectively. Element 'X' reacts with element 'Y' to form compound with molecular formula:
- (a) XY (b) XY_2
(c) XY_3 (d) X_3Y_3
22. The atomic radius (pm) of Li, Na, K and Rb varies in the order:
- (a) $\text{Na} < \text{K} < \text{Rb} < \text{Li}$ (b) $\text{K} < \text{Na} < \text{Li} < \text{Rb}$
(c) $\text{Li} < \text{Na} < \text{K} < \text{Rb}$ (d) $\text{Rb} < \text{K} < \text{Na} < \text{Li}$
23. The electron affinity value (kJ mol^{-1}) of fluorine (F) is less than:
- (a) Hydrogen (H) (b) Lithium (Li)
(c) Oxygen (O) (d) Chlorine (Cl)
24. Which of the following is an example of strong electrolyte?
- (a) H_2CO_3 (b) NH_4OH
(c) NaCl (d) $\text{COOH} - \text{COOH}$
25. If pH of a solution changes from 5 to 4, the change in hydrogen ion concentration shall be:
- (a) two times (b) five times
(c) ten times (d) twenty times
26. Bleaching powder is manufactured by the reaction of:
- (a) CaCl_2 and CaCO_3 (b) Cl_2 and Ca(OH)_2
27. (c) Cl_2 and CaSO_4 (d) Cl_2 and $\text{Ca(HCO}_3)_2$
The hard glass is obtained by fusing:
- (a) soda ash, sand and limestone
(b) a mixture of sand, lime, borax and alkali carbonates
(c) potassium carbonate and limestone
(d) potassium carbonate, lead oxide and sand
28. Which of the following metals is most reactive?
- (a) Aluminium (b) lead
(c) Mercury (d) Silver
29. Which of the following alloys contains chromium?
- (a) Steel (b) Stainless steel
(c) Magnadium (d) Brass
30. Which of the following is monomer of natural rubber?
- (a) Chloroethene (b) Chloroprene
(c) Isoprene (d) Buta-1, 3-diene
31. Alkaline KMnO_4 oxidises propanone to:
- (a) propanoic acid (b) ethanoic acid
(c) methanoic acid (d) oxalic acid
32. Which of the following metals can displace Zn from ZnSO_4 solution?
- (a) Calcium (b) Copper
(c) Iron (d) Mercury
33. Heating of sodium ethanoate with soda lime yields:
- (a) ethane (b) methane
(c) ethanol (d) methanol
34. The compound formed by the reaction of ethyne with bromine is:
- (a) $\text{Br}-\text{CH}=\text{CH}-\text{Br}$ (b) $\text{Br}-\text{CH}_2-\text{CHBr}_2$
(c) $\text{CH}_2=\text{CH}-\text{Br}$ (d) $\text{Br}_2\text{CH}-\text{CHBr}_2$
35. Cell were first discovered by Robert Hooke in the year:
- (a) 1665 (b) 1674 (c) 1831 (d) 1839
36. The plant cells have a rigid cell wall that lies:
- (a) Outside the plasma membrane
(b) Inside the plasma membrane
(c) In between the plasma membranes
(d) None of the above
37. During mitosis nucleolus and nuclear membrane are lost in stage:
- (a) Prophase (b) Metaphase
(c) Anaphase (d) Telophase
38. An example of simple and permanent tissue is:
- (a) Xylem (b) Phloem
(c) Sclerenchyma (d) All the above
39. The shape of squamous epithelial tissue is:
- (a) Cubical (b) Flattened
(c) Pillar like (d) None of the above
40. Fresh water sponge belongs to phylum:
- (a) Arthropoda (b) Annelida

Saidai Duraisamy Entrance Exam Model Question Paper

Sebastian Brünink



Saidai Duraisamy Entrance Exam Model Question Paper:

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Table of Contents Saidai Duraisamy Entrance Exam Model Question Paper

1. Understanding the eBook Saidai Duraisamy Entrance Exam Model Question Paper
 - The Rise of Digital Reading Saidai Duraisamy Entrance Exam Model Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Saidai Duraisamy Entrance Exam Model Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Saidai Duraisamy Entrance Exam Model Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Saidai Duraisamy Entrance Exam Model Question Paper
 - Personalized Recommendations
 - Saidai Duraisamy Entrance Exam Model Question Paper User Reviews and Ratings
 - Saidai Duraisamy Entrance Exam Model Question Paper and Bestseller Lists
5. Accessing Saidai Duraisamy Entrance Exam Model Question Paper Free and Paid eBooks
 - Saidai Duraisamy Entrance Exam Model Question Paper Public Domain eBooks
 - Saidai Duraisamy Entrance Exam Model Question Paper eBook Subscription Services
 - Saidai Duraisamy Entrance Exam Model Question Paper Budget-Friendly Options
6. Navigating Saidai Duraisamy Entrance Exam Model Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Saidai Duraisamy Entrance Exam Model Question Paper Compatibility with Devices
 - Saidai Duraisamy Entrance Exam Model Question Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Saidai Duraisamy Entrance Exam Model Question Paper
 - Highlighting and Note-Taking Saidai Duraisamy Entrance Exam Model Question Paper
 - Interactive Elements Saidai Duraisamy Entrance Exam Model Question Paper

8. Staying Engaged with Saidai Duraisamy Entrance Exam Model Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Saidai Duraisamy Entrance Exam Model Question Paper
9. Balancing eBooks and Physical Books Saidai Duraisamy Entrance Exam Model Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Saidai Duraisamy Entrance Exam Model Question Paper
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Saidai Duraisamy Entrance Exam Model Question Paper
 - Setting Reading Goals Saidai Duraisamy Entrance Exam Model Question Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Saidai Duraisamy Entrance Exam Model Question Paper
 - Fact-Checking eBook Content of Saidai Duraisamy Entrance Exam Model Question Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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