

Polyatomic Ions WS

Ammonium (NH_4)¹⁺
 Carbonate (CO_3)²⁻
 Chromate (CrO_4)²⁻

Dichromate (Cr_2O_7)²⁻
 Hydroxide (OH)¹⁻
 Nitrate (NO_3)¹⁻

Nitrite (NO_2)¹⁻
 Sulfate (SO_4)²⁻
 Sulfite (SO_3)²⁻

Phosphate (PO_4)³⁻

Write the **NAME** of each of the following compounds.

1. $(\text{NH}_4)\text{Cl}$ Ammonium Chloride
2. $\text{Be}(\text{SO}_4)$ Beryllium Sulfate
3. $(\text{NH}_4)_3\text{N}$ Ammonium Nitride
4. MgCl_2 Magnesium Chloride
5. $\text{NH}_4(\text{NO}_3)$ Ammonium Nitrate
6. $\text{Sr}_3(\text{PO}_4)_2$ Strontium Phosphate
7. $\text{Zn}(\text{CrO}_4)_2$ Zinc Dichromate
8. $\text{K}_2(\text{Cr}_2\text{O}_7)$ Potassium Dichromate

9. $\text{Ga}(\text{ClO}_3)_3$ Gallium Chlorate
10. $\text{Cu}(\text{OH})$ Copper I Hydroxide
11. $(\text{NH}_4)_3(\text{PO}_4)$ Ammonium Phosphate
12. $\text{Fe}(\text{SO}_4)$ Iron II Sulfate
13. $\text{Mg}(\text{NO}_2)_2$ Magnesium Nitrite
14. $(\text{NH}_4)\text{NO}_2$ Ammonium Nitrite
15. $\text{Na}_2(\text{Cr}_2\text{O}_7)$ Sodium Dichromate
16. $\text{Na}(\text{OH})$ Sodium Hydroxide

Write the **CHEMICAL FORMULA** for each of the given **NAMES**.....cross charges

- | | | | | | |
|-------------------------|---------------------------------------|---|-------------------------------|---|-----------------|
| 17. calcium carbonate | Ca^{2+} $(\text{CO}_3)^{2-}$ | = | $\text{Ca}_2(\text{CO}_3)_2$ | = | CaCO_3 |
| 18. barium nitrate | Ba^{+2} NO_3^{-1} | = | $\text{Ba}(\text{NO}_3)$ | = | |
| 19. ammonium sulfate | NH_4^{+1} SO_4^{-2} | = | $(\text{NH}_4)_2 \text{SO}_4$ | = | |
| 20. aluminum hydroxide | Al^{+3} OH^{-1} | = | $\text{Al}(\text{OH})_3$ | = | |
| 21. calcium phosphate | Ca^{+2} PO_4^{-3} | = | $\text{Ca}_3(\text{PO}_4)_2$ | = | |
| 22. cesium nitrate | Cs NO_2^{-1} | = | CsNO_3 | = | |
| 23. sodium nitrite | Na^{+1} NO_3^{-1} | = | NaNO_3 | = | |
| 24. calcium sulfate | Ca^{+2} SO_4^{-2} | = | CaSO_4 | = | |
| 25. beryllium sulfate | Be SO_4^{-2} | = | BeSO_4 | = | |
| 26. sodium carbonate | Na^{+1} CO_3^{-2} | = | Na_2CO_3 | = | |
| 27. magnesium phosphate | Mg^{+2} PO_4^{-3} | = | $\text{Mg}_3(\text{PO}_4)_2$ | = | |
| 28. calcium phosphate | Ca^{+2} PO_4^{-3} | = | $\text{Ca}_3(\text{PO}_4)_2$ | = | |

Polyatomic Ions Ws

Luigia Sabbatini, Elvira De Giglio

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located on the right side of the slide.

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