## TYPES OF CHEMICAL REACTIONS

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- Double replacement: Na<sub>2</sub>CO<sub>4</sub> + PbCl 2 NaCl + PbCO<sub>4</sub>
- 2) Single replacement: Ch + 2 NaBr -> 2 NaCl + Br<sub>2</sub>

answer key

- 3) Synthesis:  $2 H_2 + 0_2 2 HO$
- 4) Double replacement: 3 Ca(OH)<sub>2</sub> + 2 FeCl 2 Fe(OH)<sub>2</sub> + 3 CaCl<sub>2</sub>
- 5) Double replacement: AgNO<sub>3</sub> + KI -> AgI + KNO<sub>3</sub>
- 6) Single replacement: Cu + 2 AgNO<sub>3</sub> -> 2 Ag + Cu(NO<sub>3</sub>)<sub>2</sub> [Copper (II) is used here]
- 7) Decomposition: Mg(ClO<sub>3</sub>)<sub>2</sub> MgCl<sub>2</sub> + 3 0<sub>2</sub>
- 8) Single replacement: Fe + Cu(NO<sub>3</sub>)<sub>2</sub> -> Fe(NO<sub>3</sub>)<sub>2</sub> + Cu [Assume Fe<sup>2\*</sup>]
- 9) Synthesis: S+02 SO2
- 10) Decomposition: 2 Hg0 4 Hg + 0

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1) 2Mg + 0<sub>2</sub> 2 MgO Synthesis
- 2) Cal<sub>2</sub> + Cl<sub>2</sub> CaCl<sub>2</sub> In Single Replacement
- 3) 3KOH + AICl<sub>3</sub> -> AK(OH)<sub>3</sub> + 3KCl Double Replacement
- 4) C + O2 -> CO2 Synthesis

# **Reaction Types Answer Key**

Robin Donohoe Dennison, Kathleen Farrell

## **Reaction Types Answer Key:**

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#### **Table of Contents Reaction Types Answer Key**

- 1. Understanding the eBook Reaction Types Answer Key
  - The Rise of Digital Reading Reaction Types Answer Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Reaction Types Answer Key
  - $\circ \ Exploring \ Different \ Genres$
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Reaction Types Answer Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Reaction Types Answer Key
  - Personalized Recommendations
  - Reaction Types Answer Key User Reviews and Ratings
  - Reaction Types Answer Key and Bestseller Lists
- 5. Accessing Reaction Types Answer Key Free and Paid eBooks
  - Reaction Types Answer Key Public Domain eBooks
  - Reaction Types Answer Key eBook Subscription Services
  - Reaction Types Answer Key Budget-Friendly Options

- 6. Navigating Reaction Types Answer Key eBook Formats
  - o ePub, PDF, MOBI, and More
  - Reaction Types Answer Key Compatibility with Devices
  - Reaction Types Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Reaction Types Answer Key
  - Highlighting and Note-Taking Reaction Types Answer Key
  - Interactive Elements Reaction Types Answer Key
- 8. Staying Engaged with Reaction Types Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Reaction Types Answer Key
- 9. Balancing eBooks and Physical Books Reaction Types Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Reaction Types Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Reaction Types Answer Key
  - Setting Reading Goals Reaction Types Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reaction Types Answer Key
  - Fact-Checking eBook Content of Reaction Types Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

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