

Worksheet 1-1 - Measuring Reaction Rates

 Name - _____
 Due Date - _____

Answer the questions in the space provided. You must show all of your work to receive full marks. All answers MUST be rounded to the correct number of Sig. Figures.

1. A chemist wishes to determine the rate of reaction of zinc with hydrochloric acid. The equation for the reaction is:



A piece of zinc is dropped into 1.00 L of 0.100 M HCl and the following data were obtained:

Time	Mass of Zinc
0 s	0.016 g
4 s	0.014 g
8 s	0.012 g
12 s	0.010 g
16 s	0.008 g
20 s	0.006 g

$$\begin{array}{r} 0.016\text{g} \\ - 0.006\text{g} \\ \hline 0.010\text{g} \end{array}$$

- a) Calculate the *Rate of Reaction* in grams of Zn consumed per second.

$$\text{rate} = \frac{\Delta g}{\Delta \text{time}} = \frac{0.010\text{g}}{20\text{sec}}$$

Answer 5.0×10^{-4}

- b) Calculate the *Rate of Reaction* in moles of Zn consumed per second.

$$5.0 \times 10^{-4} \frac{\text{g}}{\text{sec}} \times \frac{1\text{mol}}{65.4\text{g}} =$$

Answer $7.6 \times 10^{-6} \frac{\text{mol}}{\text{sec}}$

- c) Write out the complete ionic equation for the reaction.



- d) What will happen to the $[\text{H}^+]$ as the reaction proceeds?

decrease

- e) What will happen to the $[\text{Cl}^-]$ as the reaction proceeds?

nothing, it is a spectator

Reaction Rate Practice Problems With Answers

Daniel F McAuley



Reaction Rate Practice Problems With Answers:

This book delves into Reaction Rate Practice Problems With Answers. Reaction Rate Practice Problems With Answers is an essential topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Reaction Rate Practice Problems With Answers, encompassing both the fundamentals and more intricate discussions.

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