

Calculating Average Rates of Reaction

You are required to show work for all calculations.

1. During the combustion of methane, CH_4 , shown by the reaction



The concentration of methane was measured at various time intervals and the following results were obtained:

Time (s)	$[\text{CH}_4]$ ($\text{mol} \cdot \text{l}^{-1}$)
10	2.40
20	1.20
30	0.80
40	0.60

Calculate the average rate of loss of methane during the 10 to 40 second time period.

Solution:

$$\begin{aligned}
 \text{rate} &= \frac{\Delta[\text{CH}_4]}{\Delta t} \\
 &= \frac{[0.60 - 2.4]\text{mol}}{(40 - 10)\text{s} \times \text{L}} \\
 &= \frac{-1.8\text{mol}}{30\text{s} \times \text{L}} \\
 &= -0.06 \frac{\text{mol}}{\text{L} \times \text{s}}
 \end{aligned}$$

Note: the answer is negative because we are calculating the **rate of loss of methane**.

Calculating Reaction Rates using Stoichiometry

Example:

In the following decomposition reaction,



Reaction Rate Practice Problems

Roman Wölfel

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Reaction Rate Practice Problems:

The Enigmatic Realm of **Reaction Rate Practice Problems**: Unleashing the Language is Inner Magic

In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Reaction Rate Practice Problems** a literary masterpiece penned with a renowned author, readers set about a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting impact on the hearts and minds of those that partake in its reading experience.

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