

Activity A:**Energy of chemical bonds**

Get the Gizmo ready:

- Check that **Reaction 1** and **Forward** are selected.
- Select the **INVESTIGATION** tab.



Introduction: The heat energy stored in a chemical system is called the **enthalpy** (H) of the system. When atoms are joined by a chemical bond, energy must be added to pull them apart. This increases the enthalpy of the system. When a chemical bond forms, energy is released as shared electrons move into lower-energy orbitals. This causes the enthalpy to decrease.

Question: How can you predict how much energy is released in a chemical reaction?

1. **Predict:** In the warm-up activity, you observed how the reaction inside the chamber affected the temperature of the surrounding water. Based on what happens to the surrounding water, do you think heat energy (enthalpy) is absorbed in the reaction or released? Explain.

It is released since the heat increased inside the container

2. **Observe:** In the Gizmo, the energy required to break a chemical bond is modeled by placing a molecule into a set of mechanical claws. Place one of the hydrogen (H_2) molecules between the claws, and press **Break bond**.

A. What happens? The hydrogen bond was broken as the claws pulled apart.

- B. Look under the **Energy absorbed** column of the table. How much energy was required to break this bond?

436 kJ/mol

Note: The energy is given here in units of kilojoules per mole (kJ/mol). This is the energy, in kilojoules, required to break all of the H–H bonds in one mole of H_2 gas.

- C. Remove the hydrogen atoms from the claws and then break apart the other H–H molecule. What is the total energy absorbed so far? 872 kJ/mol

3. **Measure:** Notice that the oxygen atoms are connected by a double covalent bond. This is because the oxygen atoms share two pairs of electrons. Place the oxygen molecule in the claws and press **Break bond**.

A. How much energy is required to break the first O–O bond?

349 kJ/mol

B. Press **Break bond**. How much energy is needed to break both bonds?

495 kJ/mol

Reaction Energy Modern Chemistry Answers

VM Jensen



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