

Chapter 1: REACTION KINETICS

1.1 REACTION ORDER AND RATE

- a) zero-order reaction: $\text{rate} = k$
- b) first-order reaction: $\text{rate} = k[A]$
- c) second-order reaction: $\text{rate} = k[A]^2$
- d) third-order reaction: $\text{rate} = k[A]^3$
- e) rate law: $\text{rate} = k[A]^m[B]^n$
- f) rate law: $\text{rate} = k[A]^m[B]^n$
- g) rate law: $\text{rate} = k[A]^m[B]^n$
- h) rate law: $\text{rate} = k[A]^m[B]^n$
- i) rate law: $\text{rate} = k[A]^m[B]^n$
- j) rate law: $\text{rate} = k[A]^m[B]^n$
- k) rate law: $\text{rate} = k[A]^m[B]^n$
- l) rate law: $\text{rate} = k[A]^m[B]^n$
- m) rate law: $\text{rate} = k[A]^m[B]^n$
- n) rate law: $\text{rate} = k[A]^m[B]^n$
- o) rate law: $\text{rate} = k[A]^m[B]^n$
- p) rate law: $\text{rate} = k[A]^m[B]^n$
- q) rate law: $\text{rate} = k[A]^m[B]^n$
- r) rate law: $\text{rate} = k[A]^m[B]^n$
- s) rate law: $\text{rate} = k[A]^m[B]^n$
- t) rate law: $\text{rate} = k[A]^m[B]^n$
- u) rate law: $\text{rate} = k[A]^m[B]^n$
- v) rate law: $\text{rate} = k[A]^m[B]^n$
- w) rate law: $\text{rate} = k[A]^m[B]^n$
- x) rate law: $\text{rate} = k[A]^m[B]^n$
- y) rate law: $\text{rate} = k[A]^m[B]^n$
- z) rate law: $\text{rate} = k[A]^m[B]^n$

REACTION RATE

- a) Reaction rate is the speed at which a reaction proceeds. It is defined as the change in concentration of a reactant or product per unit time.
- b) Reaction rate is defined as the change in concentration of a reactant or product per unit time.
- c) Reaction rate is defined as the change in concentration of a reactant or product per unit time.
- d) Reaction rate is defined as the change in concentration of a reactant or product per unit time.

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Reaction Energy And Reaction Kinetics Study Guide

Sebastian Brünink



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