

## CHAPTER 10: Reactions

**Reaction Energy****SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of  $\Delta H_f^\circ$  is \_\_\_\_\_.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- \_\_\_\_\_  $\Delta H_{\text{f}}^\circ$  of a substance is  $2074^\circ\text{C}$  whereas a 4.000 g of substance is heated from the decomposition for?
- \_\_\_\_\_ A substance has a specific heat of  $0.001 \text{ J/g}^\circ\text{C}$ . If a sample weighing  $10.0 \text{ g}$  is heated from  $0^\circ\text{C}$  to  $100^\circ\text{C}$ , what will be the change in enthalpy?  $100^\circ\text{C}$  is  $373^\circ\text{K}$  and  $0^\circ\text{C}$  is  $273^\circ\text{K}$ .
- The product is a chemical bond and enthalpy of  $100 \text{ kJ/mol}$ , and the substance has an enthalpy of  $100 \text{ kJ/mol}$ .
  - What is the value of  $\Delta H$  for this reaction?

# Reaction Energy Section 1 Answer Key

**RS Peters**



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