



NATIONAL
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

MARKS: 100
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use blue or black ink for writing. Do not use red ink.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fusion} for 1-hexanol in kJ mol⁻¹ and 1 mol of 1-hexanol.

Compound	Molecular mass (M _r)	Boiling point (°C)
A	102	16,2
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B or of A, B and C or of others as well? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and is the prime reason of their volatility. (2)
- 3.1.4 What factor of the compounds A and B is responsible for the difference in their boiling points? (2)
- 3.1.5 What compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 Which compound (A or B) will have the higher surface tension? (2)

Compound C (M_r 102) will show in the table, has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 16,2 °C.

- 3.1.8 Explain why the boiling point of compound C is 16,2 °C (lower than that of compound A and B) by referring to the ΔH_{vap} and intermolecular or intramolecular forces involved. (2)

3.2 Substances D and E exhibit a much stronger intermolecular forces and hence a relatively low volatility as shown in the diagram below.



- 3.2.1 Is substance D or E or both, not the better solvent? (2)
- 3.2.2 Explain the answer to QUESTION 3.2.1 about the volatility in the intermolecular forces involved. (2)

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Physical Science Paper 2 Intermolecular Forces Investigation

J.S. Rowlinson



Physical Science Paper 2 Intermolecular Forces Investigation:

Study Guide for CTET Paper 2 (Class 6 - 8 Teachers) Mathematics/ Science with Past Questions Disha

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U.S. Government Research & Development Reports ,1970 *Energy Research Abstracts* ,1990 *Physics Briefs* ,1992 *Journal of Research of the National Bureau of Standards* United States. National Bureau of Standards,1951 *Scientific and Technical Aerospace Reports* ,1995 *Nuclear Science Abstracts* ,1971 *Sir James Dewar, 1842-1923* J.S. Rowlinson,2016-03-09 Sir James Dewar was a major figure in British chemistry for around 40 years He held the posts of Jacksonian Professor of Natural Philosophy at Cambridge 1875 1923 and Fullerian Professor of Chemistry at the Royal Institution 1877 1923 and is remembered principally for his efforts to liquefy hydrogen successfully in the field that would come to be known as cryogenics His experiments in this field led him to develop the vacuum flask now more commonly known as the thermos and in 1898 he was the first person to successfully liquefy hydrogen A man of many interests he was also with Frederick Abel the inventor of explosive cordite an achievement that involved him in a major legal battle with Alfred Nobel Indeed Dewar s career saw him involved in a number of public quarrels with fellow scientists he was a fierce and sometimes unscrupulous defender of his rights and his claims to priority in a way that throws much light on the scientific spirit and practice of his day This the first scholarly biography of Dewar seeks to resurrect and reinterpret a man who was a giant of his time but is now

sadly overlooked In so doing the book will shed much new light on the scientific culture of the late nineteenth and early twentieth centuries and the development of the field of chemistry in Britain

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