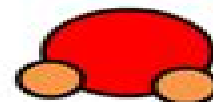


Chapter 6: Chemical Bonding Guided Notes

I. Why do Atoms Combine?

A. Chemical Formula

- The chemical formula shows:
 - The _____ in the compound.
 - The ratio of their _____



B. Chemical Bond

- Strong attractive force between _____ or ions in a molecule or compound.
- Formed by:
 - _____ electrons, (e-) by losing or gaining electrons.
 - _____ electrons (e-)

C. Stability

- Octet Rule
 - Most atoms form bonds in order to have _____ valence electrons (e-).
 - _____ outer energy level.
 - Want to be like the Noble Gases (group 18).
- Stability is the driving force behind bond formation. All atoms bond to become stable.
- Transferring e-



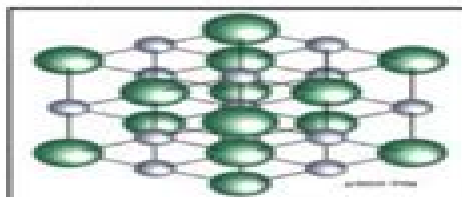
- Sharing e-



II. Kinds of Chemical Bonds

A. Ionic Bond

- Attraction between 2 _____ charged ions.
- Ions are _____ atoms.
- A _____ is a _____ charged ion. Forms when an atom loses an electron.
Formed by _____.
- A _____ is a _____ charged ion. Forms when an atom gains an electron.
Formed by _____.
- Formed by _____ electrons (e-) from a _____ to a _____.
- Ions form a 3-D crystal lattice.



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Lingsheng Yao



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Table of Contents Physical Science Chapter 6 Chemical Bonding Wordwise Answers

1. Understanding the eBook Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - The Rise of Digital Reading Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - Personalized Recommendations
 - Physical Science Chapter 6 Chemical Bonding Wordwise Answers User Reviews and Ratings
 - Physical Science Chapter 6 Chemical Bonding Wordwise Answers and Bestseller Lists
5. Accessing Physical Science Chapter 6 Chemical Bonding Wordwise Answers Free and Paid eBooks
 - Physical Science Chapter 6 Chemical Bonding Wordwise Answers Public Domain eBooks
 - Physical Science Chapter 6 Chemical Bonding Wordwise Answers eBook Subscription Services
 - Physical Science Chapter 6 Chemical Bonding Wordwise Answers Budget-Friendly Options

6. Navigating Physical Science Chapter 6 Chemical Bonding Wordwise Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Chapter 6 Chemical Bonding Wordwise Answers Compatibility with Devices
 - Physical Science Chapter 6 Chemical Bonding Wordwise Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - Highlighting and Note-Taking Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - Interactive Elements Physical Science Chapter 6 Chemical Bonding Wordwise Answers
8. Staying Engaged with Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chapter 6 Chemical Bonding Wordwise Answers
9. Balancing eBooks and Physical Books Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chapter 6 Chemical Bonding Wordwise Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - Setting Reading Goals Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - Fact-Checking eBook Content of Physical Science Chapter 6 Chemical Bonding Wordwise Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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