

SAMPLE FINAL EXAM QUESTIONS 2020

Theory - 10 questions (Final Exam will have 20 question)

Q1. We have developed a regression model: $\text{Rent} = b_0 + b_1 * \text{House_WithPool}$ where House_WithPool is a dummy variable set to 1 if a house has a Pool and is set to 0 otherwise. The Regression summary output is

	Estimate	S.E.	t Value	Pr> t
Intercept	1542.149	2.954	52.58***	<.0001
House_WithPool	253.666	3.123	81.42***	<.0001

What is the average rent of a house with Pool given the summary output?

- a. 1542.149
- b. 1795.815
- c. 1288.483
- d. 253.666

Answer: B

Q2. Why do we use Log-Transformations when we perform transformation of variables?

- A. To achieve a more linear relationship
- B. To make a distribution more normal
- C. To make the variance more constant
- D. All the above

Solution: D) A, B and C are the objectives of performing Log Transformations

Q3. What is the relation between Correlation and Covariance

$$\text{Correlation} = \frac{\text{Cov}(x, y)}{\sigma_x * \sigma_y}$$

Q4. Which of the following statement is correct regarding the "size effect"?

- A) Size effect means smaller firms have higher returns than larger firms on average
- B) Stocks with high market capitalization generate returns greater than stocks with low market capitalization
- C) Stocks with low book to market ratio tend to outperform high book to market ratio stocks
- D) Firms that invest more will have a higher stock price

Solution A: (Week 8 Lesson 1)

Q5. Which of the following is **not Fama-French factor?**

- A. Value
- B. Size
- C. Momentum
- D. Market

Answer: C) Fama-French factors include market, size, value factors and others.

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Padhraic Smyth



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American Education ,1977 **R.R. Bowker's Software for Schools** ,1987 **Tests in Education** Philip Levy,Harvey Goldstein,2014-05-19

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Education Services Curriculum Guide National Center for Educational Statistics,1975 The Transvaal Educational News ,1929 **Middle Grade Teachers' Mathematical Knowledge and Its Relationship to Instruction** Judith Sowder,Randolph A. Philipp,Barbara E. Armstrong,Bonnie P. Schappelle,1998-07-16 The outcome of a two year investigation this book shows how teachers understanding of the mathematics of number quantity and proportion influences how they teach and what their students learn of the concepts skills and reasoning associated with this mathematical domain of knowledge It grew out of the recognition of the need to understand the complexities of helping teachers reconceptualize the mathematics they teach and the resulting effects in their classrooms The book includes case studies of five teachers from different types of school settings illustrating changes in the teachers teaching methods expectations of students and beliefs about the role of professional development Journal of Personalized Instruction ,1978 **Resources in Vocational Education** ,1977 *Report of the Minister of Education* ,1909 **Reports of the Minister of Education** Ontario. Department of Education,1909 *Report of the Minister of Education* Ontario. Dept. of Education,1909 **Report of the Minister of Education** Ontario. Department of Education,1909

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow 2\text{e}^- + 2\text{H}_2\text{O} + 2\text{AuCN}$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} \rightarrow 2\text{e}^- + 2\text{H}_2\text{O} + 2\text{AuCN}$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} + 2\text{AuCN} \rightarrow 2\text{OH}^- + 2\text{H}_2\text{O} + 2\text{AuCN}$ the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions 6e^- 14H^+ $\text{Cr}_2\text{O}_7^{2-} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + \text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + 2\text{e}^-$ equalize the electrons 6e^- 14H^+ $\text{Cr}_2\text{O}_7^{2-} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O} + 6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web balance each half reaction in basic solution 4 5 6 7 $\text{Cr}_2\text{O}_7^{2-} + \text{NO}_2 \rightarrow \text{NO}_3^- + \text{MnO}_2$ $\text{Cr}_2\text{O}_7^{2-} + \text{NO}_3^- + \text{SO}_2 \rightarrow \text{Mn}_2\text{O}_3$ balance each redox reaction in acid solution using the half reaction method 8 $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} \rightarrow \text{TeO}_3 + \text{N}_2\text{O}_4$ $\text{O}_2 + \text{Cr}_3\text{Te} \rightarrow \text{ReO}_4$ 11 $\text{PbO}_2 + \text{IO}_2 \rightarrow \text{Pb}_2$ 12 as $\text{NO}_3^- + \text{IO}_3^- \rightarrow \text{Re} + \text{IO}_3^- + \text{H}_2\text{AsO}_4$

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H_2O H^+ and e^- in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ c $\text{C}_6\text{H}_6\text{O}_6$ and NO_2 d $\text{C}_6\text{H}_6\text{O}_6$ and NO

2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h and

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