Balancing Act

Name

Atoms are not <u>created</u> or <u>destroyed</u> during a chemical reaction.

Scientists know that there must be the <u>same</u> number of atoms on each <u>side</u> of the <u>reaction</u>. To balance the chemical equation, you must add <u>coefficients</u> in front of the chemical formulas in the equation. You cannot <u>add</u> or <u>remove</u> subscripts.

- 1) Determine number of atoms for each element
- 2) Pick an element that is not equal on both sides of the equation.
- 3) Add coefficients in front of the element and adjust your counts.
- Continue adding coefficients until you get the same number of elements on each side.

Example:

$$2$$
 Mg + O_2 \rightarrow 2 MgO
Mg = 2 Mg = 2
 $O = 2$ $O = 2$

Try these:

$$N_2$$
 + 3 H_2 \rightarrow 2 NH_3
 $N = 2$
 $H = 6$ $H = 6$

$$2 H_2O_2 \rightarrow 2 H_2O + O_2$$
 $H = 4 H = 4$
 $O = 4 O = 4$

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KJ Lindholm-Leary

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Balancing Act Julie Plummer, 2022 **Balancing Act** Paula Skelley,1995 Balancing Act Dana Allin, 2010 The Balancing Act Richard Cavendish, 1960 Zuo shou de zhang wen ,2003 Balancing Act John Berbrich, 2003 Balancing Act Rochelle Ratner, 2006 The Balancing Act Eileen Gloster, 1994 Balancing Act Spain & Bianchi, 2004 Balancing Act Fred Guzzi, 2021 **Balancing Act** Ralph N. Clough,1987-09 Balancing Act Pomegranate Communications, Inc. Balancing Act Dana Allin, 2010

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