



CBSE



SAMPLE QUESTION PAPER SOCIAL SCIENCE (087) CLASS-X (2023-24)

Time Allowed: 3 Hours

Maximum Marks: 80

General Instructions:

1. The question paper comprises Six Sections – A, B, C, D, E and F. There are 37 questions in the Question paper. All questions are compulsory.
2. **Section A** – From questions 1 to 20 are MCQs of 1 mark each.
3. **Section B** – Question no. 21 to 24 are Very Short Answer Type Questions, carrying 2 marks each. Answer to each question should not exceed 40 words.
4. **Section C** contains Q.25 to Q.29 are Short Answer Type Questions, carrying 3 marks each. Answer to each question should not exceed 60 words.
5. **Section D** – Question no. 30 to 33 are long answer type questions, carrying 5 marks each. Answer to each question should not exceed 120 words.
6. **Section-E** - Questions no from 34 to 36 are case based questions with three sub questions and are of 4 marks each. Answer to each question should not exceed 100 words.
7. **Section F** – Question no. 37 is map based, carrying 5 marks with two parts, 37a from History (2 marks) and 37b from Geography (3 marks).
8. There is no overall choice in the question paper. However, an internal choice has been provided in few questions. Only one of the choices in such questions have to be attempted.
9. In addition to this, separate instructions are given with each section and question, wherever necessary.
10. **Note:** CBQ stands for "Competency Based Question". 50% weightage allocated for competency-based questions.

SECTION A		
MCQs (1X20=20)		
1.	Identify the correct option that describes the act given below. i. The Act was passed by the Imperial Legislative Council. ii. It gave power to the government to repress political activities. iii. It empowered the government to detain political prisoners without trial.	1

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web in this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces metallic copper which is seen precipitating as a finely divided red powder

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web in this lab we will use experimental evidence and the principles of stoichiometry to determine which ion is formed in a single replacement reaction between iron and copper ii sulfate pre lab assignment in preparation for the lab activity we will investigate another single replacement reaction between copper and silver nitrate

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of

the reaction between 7.0002g away copper ii sulfate and 2.0101g iron metal which produced 2.3037g of copper product with a percent yield of 100.72 introduction

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web stoichiometry lab report the reaction of iron with copper ii sulfate period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding heated CuSO_4 solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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