



KWAZULU-NATAL PROVINCE
EDUCATION
REPUBLIC OF SOUTH AFRICA

**NATIONAL
SENIOR CERTIFICATE**

GRADE 11

PHYSICAL SCIENCES: PHYSICS (P1)

NOVEMBER 2023

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$$8 3 \text{ mole h}_2\text{s} \times \frac{2 \text{ mole h}_2\text{o}}{2 \text{ mole h}_2\text{s}} = 8 3 \text{ mole h}_2\text{o}$$

 8 3 mole h₂s x 18g h₂o i mole h₂o 149 4g h₂o is your theoretical yield so percent yield 137 1g h₂o actual yield 149 4g h₂o x 100
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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

 if the reaction of 21 3 grams of o₂ produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have

$$\frac{21 3 \text{ g O}_2}{32 \text{ g/mol}} = 0 666 \text{ mol O}_2$$

$$\frac{21 3 \text{ g O}_2}{32 \text{ g/mol}} \times \frac{3 \text{ mol CO}}{3 \text{ mol O}_2} = 0 666 \text{ mol CO}$$

$$0 666 \text{ mol CO} \times 28 \text{ g/mol} = 18 6 \text{ g CO}$$

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 calculated mass of the product expected x100 in this reaction a yield 80 is anticipated

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 $80\text{ g I}_2\text{O}_5 \times \frac{1\text{ mol I}_2\text{O}_5}{1\text{ mol I}_2} \times \frac{1\text{ mol I}_2}{1\text{ mol I}_2\text{O}_5} = 1.333\text{ mol I}_2$
 $28\text{ g CO} \times \frac{1\text{ mol CO}}{28\text{ g CO}} = 1\text{ mol CO}$

7.3 limiting reactant and percent yield problems - Jun 13 2023

web Jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage
 $\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100$ label 3.7.3