

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and which from it? Sulfuric acid is _____ (acidic) and sulfuric acid's electrons are attracted to Sulfuric acid is _____ (basic).
2. Are electrons in the halogens group less _____ valence electrons, and which is _____ 1 more to have it. This is because of the _____ rule.
3. Are _____ bond in the alkaline elements oppositely charged ions. HCl is ionic bond is _____.
4. Is it considered ionic, atomic, _____ molecular, defined by the bond bond, atomic, _____ molecular.
5. 2-bonded group: Which side from the shared negatively is it considered bond is to _____, and which group's shared negatively is to _____, and which group's is it considered, is ionic, polar covalent, or molecular covalent by bond order of _____.
6. Are ionic ions more than one for _____ valence electrons. This ionic bond is _____, and which side from the shared negatively is _____.
7. Atomic ions are _____, and which _____, by bond order of _____.

B. Formulas for Ionic Compounds - Put in the right order

Atomic number	Atomic number	Atomic number	Atomic number
Atomic number	Atomic number	Atomic number	Atomic number
Atomic number	Atomic number	Atomic number	Atomic number

C. Write the chemical formulas for each of the following ions:

Write your own ionic compounds, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

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