

PHYSICAL CHEMISTRY

(Complete Syllabus)

SECTION - A

Straight Objective Type

This section contains 220 multiple choice questions. Each question has 4 choices (1), (2), (3) and (4), out of which **ONLY ONE** is correct.

Choose the correct answer :

- The mass of oxygen required to produce enough CO which completely reduces x gm of Fe_2O_3 . Suppose molecular weight of $\text{O}_2 = y$, molecular weight of $\text{Fe}_2\text{O}_3 = z$
(1) $\frac{2xz}{3y}$ (2) $\frac{3xy}{2z}$
(3) $\frac{2y}{3xz}$ (4) $\frac{3z}{2xy}$
- The number of lone pair of electrons present in 2.56 gm of orthorhombic sulphur will be
(1) $0.02 N_0$ (2) $0.16 N_0$
(3) $0.32 N_0$ (4) $0.08 N_0$
- A 10.0 g of sample of mixture of calcium chloride and sodium chloride is treated with Na_2CO_3 to precipitate the calcium as calcium carbonate. This CaCO_3 is heated to convert all calcium to CaO and final mass of CaO is 1.62 gms. The % by mass of CaCl_2 in the original mixture is
(1) 15.2%
(2) 32.1%
(3) 21.8%
(4) 11.07%
- 1000 ml of oxygen at STP were passed through an ozoniser and the resulting volume was 888 ml at STP. This quantity of ozonised oxygen is passed through excess of KI solution. The weight of I_2 liberated is
(1) 25.4 g (2) 12.7 g
(3) 2.54 g (4) 1.27 g
- x_1 ml solution having normality 1 is diluted to get a solution of normality y_2 . What would be volume required for dilution?
(1) $x_1 \left(1 - \frac{y_1}{y_2}\right)$ (2) $x_1 \left(\frac{y_2}{y_1} - 1\right)$
(3) $x_1 \left(\frac{y_1}{y_2} - 1\right)$ (4) $x_1 \left(1 + \frac{y_1}{y_2}\right)$
- 1.625 g of an interhalogen compound ICl_x was subjected to electric discharge which split it into constituents and brought to STP. It occupied 112 ml. The interhalogen compound is
(1) ICl (2) ICl_3
(3) ICl_5 (4) ICl_7
- If 12.25 g of a sample of KClO_3 of 75% purity and rest vapourisable impurities is strongly heated to give KCl , KClO_4 and 0.32 gm O_2 . The mass of residue obtained is
(1) 1.769 g (2) 7.096 g
(3) 5.327 g (4) 8.85 g
- A 2 g sample of xenon reacts with fluorine. The mass of compound produced is 3.158 g. The empirical formula of the compound is
(1) XeF_2 (2) XeF_4
(3) XeF_6 (4) XeF_8
- 1 g of carbonate of metal was dissolved in 25 ml of 1 N HCl. The resulting liquid required 5 ml of 1 N NaOH for neutralisation. The equivalent weight

Physical Chemistry Guide University Questions And Answers

Jayashree Ghosh



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Laboratory Guide to Biochemistry, Enzymology, and Protein Physical Chemistry Marc le Maire, Raymond Chabaud, Guy Hervé, 2012-12-06 The study of a single well chosen substance here aspartate transcarbamylase can provide an excellent basis for a laboratory course The student is introduced to a variety of scientific ideas and to many experimental and interpretive techniques This enzyme is readily available is relatively stable has an extensive literature and its behavior has many facets substrate inhibition a large change in structure upon homotropic activation by substrates allosteric stimulation by ATP allosteric inhibition by CTP synergistic with VTP positive cooperativity for substrates negative cooperativity for CTP binding and dissociation and reassembly of subunits C and R2 from the holoenzyme CI 5 In addition 3 6 to the known biochemical aspects of these properties the results obtained here can be interpreted in the light of the high resolution X ray diffraction structures of the T and R forms the low angle X ray scattering results and the large number of mutants now available by recombinant DNA methods Future development of this course could also involve part of these methods as well as the carefully chosen experiments described here This approach resembles research more than the approaches one usually finds in biochemical laboratory courses A consistent development of ideas about a single enzyme which shows so many facets in its behavior is sure to hold the interest of the student Moreover one explores a depth and reasons to move forward that are an essential part of research

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