

| Question number | Answer                                                                                                                                                                                                                                                                                                                                  | Marks                | Guidance |
|-----------------|-----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------|----------|
| 1               | $\text{Ca} : \text{N} : \text{O} = 30.35/40.1 : 21.20/14.0 : 48.45/16.0$<br>$= 0.7569 : 1.5286 : 3.028$<br>Formula = $\text{CaN}_2\text{O}_4$                                                                                                                                                                                           | B1<br>B1             |          |
| 2 (a)           | $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$                                                                                                                                                                                                                                                    | B1                   |          |
| 2 (b)           | $4.50 \times 10^{-3} \text{ mol}$                                                                                                                                                                                                                                                                                                       | B1                   |          |
| 2 (c)           | $n(\text{KClO}_3) = 3.0 \times 10^{-3} \text{ mol}$<br>$M(\text{KClO}_3) = 122.6 \text{ g mol}^{-1}$<br>mass of $\text{KClO}_3 = 0.3678 \text{ g}$                                                                                                                                                                                      | B1<br>B1<br>B1       |          |
| 3               | $M(\text{ZnSO}_4) = 161.5 \text{ g mol}^{-1}$<br>$n(\text{ZnSO}_4) = 6.57 \times 10^{-3} \text{ mol}$<br>$n(\text{H}_2\text{O}) = (1.893 - 1.061)/18 = 4.62 \times 10^{-2} \text{ mol}$<br>$x = n(\text{H}_2\text{O})/n(\text{ZnSO}_4) = 7$<br>Formula = $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$                                      | B1<br>B1<br>B1<br>B1 |          |
| 4               | $M(\text{Na}_2\text{CO}_3) = 106.0 \text{ g mol}^{-1}$<br>Actual $n(\text{Na}_2\text{CO}_3) = 0.0234 \text{ mol}$<br>Theoretical $n(\text{Na}_2\text{CO}_3) = 0.0234 \times 100/65$<br>$= 0.0360 \text{ mol}$<br>Theoretical $n(\text{NaHCO}_3) = 0.0720 \text{ mol}$<br>Mass of $\text{NaHCO}_3 = 0.0720 \times 84.0 = 6.05 \text{ g}$ | B1<br>B1<br>B1<br>B1 |          |
| 5 (a)           | $\text{Fe}_2\text{O}_3(\text{s}) + 3\text{CO}(\text{g}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{CO}_2(\text{g})$                                                                                                                                                                                                                      | B1                   |          |
| 5 (b)           | $M(\text{Fe}_2\text{O}_3) = 159.6 \text{ g mol}^{-1}$<br>$n(\text{Fe}_2\text{O}_3) = (10\,000 \times 10^6)/159.6 = 6.266 \times 10^7 \text{ mol}$<br>$n(\text{Fe}) = 1.253 \times 10^8 \text{ mol} = 6.992 \times 10^5 \text{ g}$<br>(6992 tonnes)                                                                                      | B1<br>B1<br>B1       |          |
| 6 (a)           | $\text{Na} : \text{N} : \text{O} = 27.1/23.0 : 16.5/14.0 : 56.4/16.0$<br>$= 1.178 : 1.179 : 3.525$<br>Formula = $\text{NaNO}_3$                                                                                                                                                                                                         | B1<br>B1             |          |
| 6 (b)           | $2\text{NaNO}_3(\text{s}) \rightarrow 2\text{NaNO}_2(\text{s}) + \text{O}_2(\text{g})$                                                                                                                                                                                                                                                  | B1                   |          |

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corporation and was instrumental in popularizing chocolate candy throughout much of the world following an incomplete  
rural school education hershey was apprenticed at age 15 to a confectioner in

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battle of gettysburg from his home the two major streets in hershey pennsylvania are cocoa avenue and chocolate avenue  
during world war ii hershey made special ration bars for the troops called field ration d bars his

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influence in his life she raised hershey in the discipline of mennonite faith which taught him perseverance and success his  
father was a dream chaser and traveled in pursuit of business success

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