Chemistry--Unit 10: Chemical Periodicity
Practice Problems

I. Classification of the elements

1) Use the periodic table to write the electron configuration for silicon.

$1s^2 2s^2 2p^6 3s^2 3p^2$ or abbreviated form*: [Ne] $3s^2 3p^2$ *may use this form now!!!

2) Use the periodic table to write the electron configuration for iodine.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^5$ or [Kr] $5s^2 4d^{10} 5p^5$

3) Which family of elements is characterized by an s^2p^3 configuration?

Nitrogen Family, or Group 15

4) Name the element that matches the following description:
a. one that has 5 outer electrons on the third period of the periodic table P, phosphorus
b. one with a $4s^2 4p^5$ electron configuration Br, bromine
c. the 4th period Group 16 element Se, selenium

5) Identify the elements that have the following outer electron configurations?
a. $2s^2 2p^4$ O, oxygen
b. $4s^2$ Ca, Calcium
c. $4s^2 3d^{10}$ Zn, zinc

6) What is the common characteristic of the electron configurations of the elements Ne and Ar? In which Group would you find them?

full outermost energy level, Group 18, the noble or inert gases

7) Why would you expect lithium (Li) and sulfur (S) to have different chemical and physical properties?

they are in different groups on the periodic table (Li group 1, S group 16)

8) What characterizes silver (Ag) and iron (Fe) as transition metals?

Their electron configurations both end in the d sublevel (last electrons added are to a d sublevel).

II. Periodic Trends

9) Explain why a magnesium atom is smaller than atoms of both sodium and calcium.

Na has less +/– attraction; Ca has another energy level

10) Predict the size of the astatine (At) atom compared to that of tellurium (Te).

At would be larger than Te due to the extra energy level.

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Table of Contents Prentice Hall Chemistry 3 1 Practice Problems Answers

1. Understanding the eBook Prentice Hall Chemistry 3 1 Practice Problems Answers
    - The Rise of Digital Reading Prentice Hall Chemistry 3 1 Practice Problems Answers
    - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry 3 1 Practice Problems Answers
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an Prentice Hall Chemistry 3 1 Practice Problems Answers
    - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry 3 1 Practice Problems Answers
    - Personalized Recommendations

- Prentice Hall Chemistry 3 1 Practice Problems Answers User Reviews and Ratings
- Prentice Hall Chemistry 3 1 Practice Problems Answers and Bestseller Lists

5. Accessing Prentice Hall Chemistry 3 1 Practice Problems Answers Free and Paid eBooks
   - Prentice Hall Chemistry 3 1 Practice Problems Answers Public Domain eBooks
   - Prentice Hall Chemistry 3 1 Practice Problems Answers eBook Subscription Services
   - Prentice Hall Chemistry 3 1 Practice Problems Answers Budget-Friendly Options
6. Navigating Prentice Hall Chemistry 3 1 Practice Problems Answers eBook Formats
   - ePub, PDF, MOBI, and More
   - Prentice Hall Chemistry 3 1 Practice Problems Answers Compatibility with Devices
   - Prentice Hall Chemistry 3 1 Practice Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
   - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry 3 1 Practice Problems Answers
   - Highlighting and Note-Taking Prentice Hall Chemistry 3 1 Practice Problems Answers
   - Interactive Elements Prentice Hall Chemistry 3 1 Practice Problems Answers
8. Staying Engaged with Prentice Hall Chemistry 3 1 Practice Problems Answers
   - Joining Online Reading Communities
   - Participating in Virtual Book Clubs
   - Following Authors and Publishers Prentice Hall Chemistry 3 1 Practice Problems Answers
9. Balancing eBooks and Physical Books Prentice Hall Chemistry 3 1 Practice Problems Answers
   - Benefits of a Digital Library
   - Creating a Diverse Reading Collection Prentice Hall Chemistry 3 1 Practice Problems Answers
10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry 3 1 Practice Problems Answers
    - Setting Reading Goals Prentice Hall Chemistry 3 1 Practice Problems Answers
    - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry 3 1 Practice Problems Answers
    - Fact-Checking eBook Content of Prentice Hall Chemistry 3 1 Practice Problems Answers

- Distinguishing Credible Sources

13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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