

Algebra 1 Chapter 1

PRACTICE TEST ANSWER KEY

1) Evaluate each expression if $a = 5$, $b = 2$, $c = 2$.

(a) $2a$ **10** (b) $(2c) - 1$ **3** (c) $(2a + 3b) + c$ **8**

2) Evaluate each expression in BOLD for the given values of the variables.

(a) Perimeter of a triangle: **$(a + b) + c$** if $a = 4$, $b = 5$, $c = 10$ **19**

(b) Area of the rectangle: **lw** if $l = 16$ and $w = 2$ **32**

3) Simplify each expression.

(a) $8 + 7 \cdot 3$ **29** (b) $(8 + 7)3$ **45** (c) $\frac{6(5 - 4)}{3 \cdot 5 - 3}$ **$\frac{1}{2}$**

4) Evaluate each expression if $t = 4$, $x = 2$, $y = 1$, and $z = 6$.

(a) $x(y + z)$ **14** (b) $zt - y$ **23** (c) $\frac{(zt) - 6}{8 - x \cdot y}$ **3**

5) Solve each equation over the domain $\{0, 1, 2, 3, 4, 5, 6, 7\}$

(a) $4c - 10 = 10$ **{5}** (b) $7p = p \cdot 7$ **{0, 1, 2, 3, 4, 5, 6, 7}** (c) $43 = 8g - 5$ **{6}**

6) Translate each phrase into a variable expression. Use n as the variable.

(a) The quotient of 3 and a number **$\frac{n}{3}$** (b) Three-fourths of a number **$\frac{3}{4}n$**
 (c) 1 less than a number **$n - 1$** (d) 6 more than twice a number. **$2n + 6$**

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-}$ Fe^{2+} Cr^{3+} Fe^{3+} solution 1 balanced half reactions 6e^- 14H^+ $\text{Cr}_2\text{O}_7^{2-}$ 2Cr^{3+} $7\text{H}_2\text{O}$ Fe^{2+} Fe^{3+} e 2 equalize the electrons 6e^- 14H^+ $\text{Cr}_2\text{O}_7^{2-}$ 2Cr^{3+} $7\text{H}_2\text{O}$ 6Fe^{2+} 6Fe^{3+} 6e^- multiplied by a factor of 6

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H^+ H_2O OH^- and e^- in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6$ 2HNO_2 $\text{C}_6\text{H}_6\text{O}_6$ $2\text{H}_2\text{O}$ 2NO choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 no 2 and $\text{C}_6\text{H}_8\text{O}_6$ b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ h and NO_2 c h and