

		<p>gerangskik ☑ is om die hoeke tussen hulle te maksimeer. ☑ Strukturele elektronpare is bindingspare plus alleenpare.</p> <p>OF</p> <p>Valensdop-elektronpaarafstotingsmodel is 'n model om die vorms van molekules te voorspel ☑ waarin die elektronpaar vanaf die buitenste dop ☑ van 'n verwysingsatoom rondom hierdie atoom gerangskik is om die afstoting tussen hulle te minimaliseer. ☑</p> <p>3 merks/punte</p>	
1.2.4		<p>The molecule is a tetrahedron ☑ Die molekule is 'n tetraëder ☑</p> <p>1 merk/punt</p>	LO2
1.2.5		 <p>Name trigonal pyramid ☑ Noem trigonale piramidele</p> <p>3 merks/punte</p>	LO4
1.3	1.3.1	<p>The electron pair is equally shared between the two Cl atoms ☑ and the resulting Cl₂ molecule has no degree of polarity. ☑ Die elektronpaar word ewe veel gedeel tussen die twee Cl-atome ☑ en die resulterende Cl₂-molekule het geen mate van polariteit nie ☑.</p> <p>2 merks/punte</p>	LO2
	1.3.2	<p>The tendency ☑ of an atom to pull the bonded pair ☑ of electrons towards itself. ☑ Die neiging ☑ van 'n atoom om die gebinde elektronpare ☑ na homself te trek ☑.</p> <p>3 merks/punte</p>	LO2
	1.3.3	<p>Electronegativity of Na = 6 Electronegativity of Cl = 3 Difference is 3 ☑</p> <p>With a difference of greater than 2.1 ☑ electron transfer will take place and the bond ☑ is ionic. ☑ (value of 2.1 does not need to be given for this mark) Elektronegatiwiteit van Na = 6 Elektronegatiwiteit van Cl = 3 Verskil is 3 ☑</p> <p>Met 'n verskil van meer as 2,1 ☑ sal elektroneoordrag plaasvind en die binding is ionies. ☑ (waarde van 2.1 hoef nie vir hierdie punt gegee te word nie)</p> <p>4 merks/punte</p>	LO2

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