

Reversible Reactions and Chemical Equilibrium Worksheet

Supporting questions

- 1) Can you draw the reversible arrow symbol?
- 2) What do you observe when you heat up hydrated copper sulfate?
- 3) Define the term dynamic equilibrium?
- 4) Why can dynamic equilibrium only occur in a closed system?
- 5) What are the 3 factors that alter the position of equilibrium?
- 6) What happens to the position of equilibrium if you add a catalyst?
- 7) For the chemical reaction below can you say what happens to the concentration of ammonia if you (a) Add in more hydrogen (b) increase the pressure (c) increase the temperature (the forward reaction is **exothermic**).



Nitrogen Hydrogen Ammonia

- 8) Why might a high temperature still be used?

Thanks for taking the time to complete these questions, you can check your answers on the next page.

Reversible Reactions And Equilibrium Ch 18 2

Worksheet

M Mosston



Reversible Reactions And Equilibrium Ch 18 2 Worksheet:

Embracing the Tune of Term: An Emotional Symphony within **Reversible Reactions And Equilibrium Ch 18 2 Worksheet**

In a world consumed by screens and the ceaseless chatter of quick connection, the melodic beauty and mental symphony created by the prepared word often disappear into the backdrop, eclipsed by the persistent noise and interruptions that permeate our lives. Nevertheless, located within the pages of **Reversible Reactions And Equilibrium Ch 18 2 Worksheet** a charming fictional treasure overflowing with raw thoughts, lies an immersive symphony waiting to be embraced. Crafted by a masterful composer of language, this charming masterpiece conducts readers on an emotional trip, well unraveling the concealed melodies and profound affect resonating within each cautiously constructed phrase. Within the depths of this touching review, we can explore the book's main harmonies, analyze its enthralling publishing style, and submit ourselves to the profound resonance that echoes in the depths of readers' souls.

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