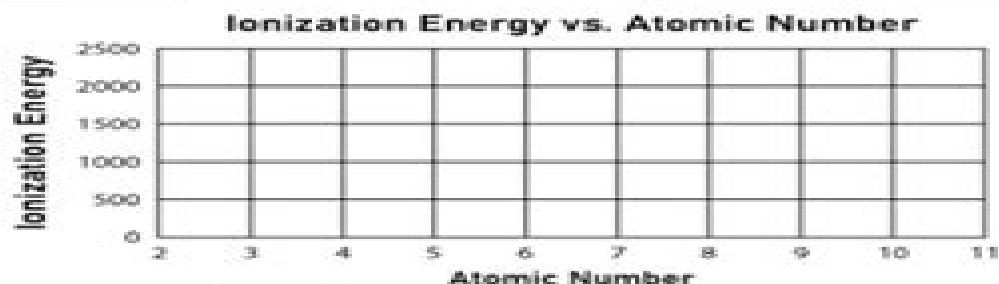


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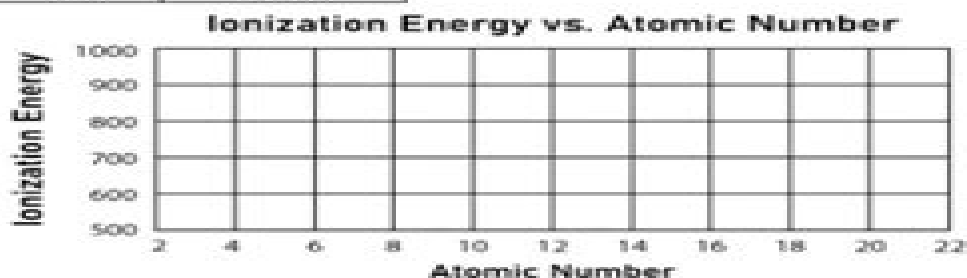
1. Using the ionization energy (IE) values of the elements in Period 2 listed below, make a line graph of IE vs. the atomic number.

Period 2						
Element	Atomic Number	IE (kJ/mol)		Element	Atomic Number	IE (kJ/mol)
Li	3	519		Li	3	519
Be	4	900		Be	4	900
B	5	799		B	5	799
C	6	1088		C	6	1088



- a) What trend do you notice for the ionization energies in Period 2?
- b) Explain why you observe this trend.
2. Make a scatter plot of ionization energy vs. atomic number for the first three atoms in Group 2A listed below.

Group 2A		
Element	Atomic Number	IE (kJ/mol)
Be	4	900
Mg	12	736
Ca	20	590



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