

Chemistry - Review Chapters (415) 6

Directions: Answer each of the following questions. You need not use complete sentences.

1. By what property did Mendeleev arrange the elements?

ATOMIC MASS

2. By what property did Moseley suggest that the periodic table be arranged?

ATOMIC NUMBER

3. What is a period? How many are there in the periodic table?

A HORIZONTAL ROW IN THE PERIODIC TABLE; 7

4. What is a group (also called a family)? How many are there in the periodic table?

A VERTICAL COLUMN IN THE PERIODIC TABLE; 18

5. State the number of valence electrons in an atom of:

a. sulfur 6	b. calcium 2	c. chlorine 7	d. arsenic 5
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6. Give the names and chemical symbols for the elements that correspond to these atomic numbers:

a. 10 Ne, NEON	b. 18 Ar, ARGON	c. 36 Kr, KRYPTON	d. 90 Th, THORIUM
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7. List, by number, both the period and group of each of these elements.

	<u>Symbol</u>	<u>Period</u>	<u>Group</u>
a. beryllium	Be	2	2
b. iron	Fe	4	8
c. lead	Pb	6	14

8. Which of the following pairs of elements belong to the same period?

<u>a. Na and Cl</u>	b. Na and Li	c. Na and Cu	d. Na and Ne
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9. Which of the following pairs of elements belong to the same group?

a. H and He	b. Li and Be	<u>c. C and Pb</u>	d. Ga and Ge
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10. How does an element's period number relate to the number of the energy level of its valence electrons?

PERIOD NUMBER = ENERGY LEVEL OF VALENCE ELECTRONS

11. What are the transition elements?

GROUPS 3-12

12. In what type of orbitals are the actinide and lanthanide electrons found?

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Table of Contents Prentice Hall Chemistry Practice Test Answers

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 - The Rise of Digital Reading Prentice Hall Chemistry Practice Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Practice Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Practice Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Practice Test Answers
 - Personalized Recommendations
 - Prentice Hall Chemistry Practice Test Answers User Reviews and Ratings
 - Prentice Hall Chemistry Practice Test Answers and Bestseller Lists
5. Accessing Prentice Hall Chemistry Practice Test Answers Free and Paid eBooks

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 - Prentice Hall Chemistry Practice Test Answers eBook Subscription Services
 - Prentice Hall Chemistry Practice Test Answers Budget-Friendly Options
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- ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Practice Test Answers Compatibility with Devices
 - Prentice Hall Chemistry Practice Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Practice Test Answers
 - Highlighting and Note-Taking Prentice Hall Chemistry Practice Test Answers
 - Interactive Elements Prentice Hall Chemistry Practice Test Answers
8. Staying Engaged with Prentice Hall Chemistry Practice Test Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Practice Test Answers
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Practice Test Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Practice Test Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Practice Test Answers
- Setting Reading Goals Prentice Hall Chemistry Practice Test Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Practice Test Answers
- Fact-Checking eBook Content of Prentice Hall Chemistry Practice Test Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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