

14

THE BEHAVIOR OF GASES

Chapter Test II

A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct answer next to the item.

Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the ratio of the change in volume to the change in temperature).
2. The volume of a fixed amount of a gas at constant pressure in the standard state (0°C and 1 atm).
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used.
4. An experimental gas constant determined from the ideal gas law at low pressures.
5. A constant of nature that has values of about 8.314 J/K·mol.
6. The a parameter of gas constant temperatures for determining the gas constant from the ideal gas law.
7. An experimental gas constant determined from the ideal gas law at low pressures.
8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. The amount of heat enough to raise 1 g of water 1°C.
10. Ratio of volume to mass

Column II

- a. standard gas law
- b. ideal gas constant R
- c. density
- d. compressibility
- e. Boyle's law
- f. molar volume
- g. standard molar volume
- h. volume
- i. compressibility
- j. heat of fusion of water

B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. Subliming the solvent of a solid sample for mass determination (weighing) requires a constant pressure of _____.
- a. the atmosphere (1 atm)
- b. the vacuum of a hard vacuum
- c. the pressure of the water
- d. the vacuum of a hard vacuum

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Arturo Cuomo



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