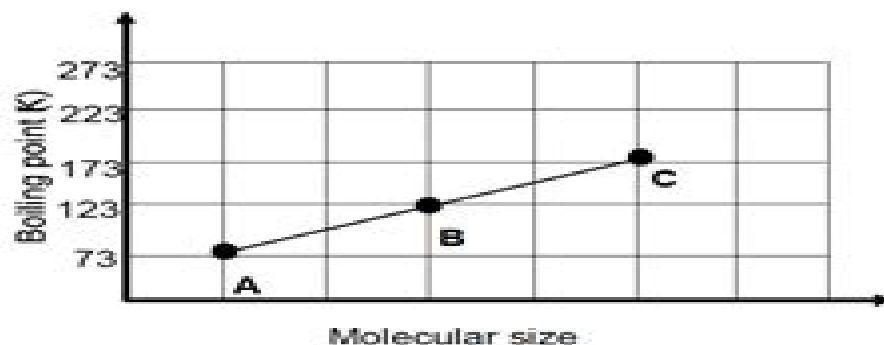


QUESTION 2

2.1 The graph of molecular size versus the boiling point is given below.

The letters **A**, **B** and **C** represent the compounds CH_4 , C_2H_6 and C_3H_8 respectively.



2.1.1 Define the term *boiling point*. (2)

2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)

2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)

2.2 Consider the two molecules in the table below.

NAME OF SUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT ($^{\circ}\text{C}$)
Ammonia	NH_3	17	- 33
Phosphine	PH_3	x	- 87,4

2.2.1 Determine the molecular mass (x) of phosphine (1)

2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

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1. Understanding the eBook Physical Science Grade 1term Test Question Paper 2caps
 - The Rise of Digital Reading Physical Science Grade 1term Test Question Paper 2caps
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Grade 1term Test Question Paper 2caps
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Grade 1term Test Question Paper 2caps
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Grade 1term Test Question Paper 2caps
 - Personalized Recommendations
 - Physical Science Grade 1term Test Question Paper 2caps User Reviews and Ratings
 - Physical Science Grade 1term Test Question Paper 2caps and Bestseller Lists
5. Accessing Physical Science Grade 1term Test Question Paper 2caps Free and Paid eBooks

- Physical Science Grade 1term Test Question Paper 2caps Public Domain eBooks
 - Physical Science Grade 1term Test Question Paper 2caps eBook Subscription Services
 - Physical Science Grade 1term Test Question Paper 2caps Budget-Friendly Options
6. Navigating Physical Science Grade 1term Test Question Paper 2caps eBook Formats
- ePub, PDF, MOBI, and More
 - Physical Science Grade 1term Test Question Paper 2caps Compatibility with Devices
 - Physical Science Grade 1term Test Question Paper 2caps Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Physical Science Grade 1term Test Question Paper 2caps
 - Highlighting and Note-Taking Physical Science Grade 1term Test Question Paper 2caps
 - Interactive Elements Physical Science Grade 1term Test Question Paper 2caps
8. Staying Engaged with Physical Science Grade 1term Test Question Paper 2caps
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Grade 1term Test Question Paper 2caps
9. Balancing eBooks and Physical Books Physical Science Grade 1term Test Question Paper 2caps
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Grade 1term Test Question Paper 2caps
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Grade 1term Test Question Paper 2caps
- Setting Reading Goals Physical Science Grade 1term Test Question Paper 2caps
 - Carving Out Dedicated Reading Time
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- Fact-Checking eBook Content of Physical Science Grade 1term Test Question Paper 2caps
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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