#### PHYSICAL SCIENCES - GRADE 10

#### CHAPTER - 1

## Chemical Reactions and Equations

- q Chemical Reaction: —Whenever a chemical change occurs we can say that a chemical reaction has taken place
  - eg Food gets digested in our body
    - Rusting of iron.
- q Chemical Equation :—A chemical reaction can be expressed symbolically by using chemical equation

eg magnesium is burnt into air to form magnesium oxide can be represented as

$$Mg + O_2 \rightarrow MgO$$

- We can observe or recognise a chemical reaction by observing change in state, colour, by evolution of gas or by change in temperature.
- Physical state of the reactant and products are mentioned to make chemical reaction more informative, eg we use (g) for gas, (l) for liquid, (s) for solid and (aq) for aqueous.
- q Balancing Equation :— We balance the chemical equation so that no. of atoms of each element involved in the reaction remain same at the reactant and product side.

eg Fe + 
$$H_2O \rightarrow Fe_2O_3 + H_2$$
 can be written as  
 $3 Fe(s) + 4H_2O(g) \rightarrow Fe_2O_3(s) + 4H_4(g)$ 

 Combination Reaction :- The reaction in which two or more substances combine to form a new single substance

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**Annelies Wilder-Smith** 

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