

Stoichiometry Worksheet

I. ONE REACTANT IN EXCESS – COMBUSTION

You burn an excess of butane (C₄H₁₀) in a limited amount of oxygen (O₂). If 67.2 L of CO₂ gas is formed, what volume of oxygen gas was present at the start of the reaction?

STEP 1: Balance the reaction:



STEP 2: Use a molar conversion to convert quantity of product given to moles:

$$67.2 \text{ L CO}_2 * \frac{(1 \text{ mol CO}_2)}{(22.4 \text{ L CO}_2)} = 3 \text{ mol CO}_2$$

STEP 3: Use a molar ratio to calculate the moles of reactant that I'm asking about:

$$3 \text{ mol CO}_2 * \frac{(13 \text{ mol O}_2)}{(8 \text{ mol CO}_2)} = 4.875 \text{ mol O}_2$$

STEP 4: Use a molar conversion to convert moles of reactant to the units I ask for:

$$4.875 \text{ mol O}_2 \frac{(22.4 \text{ L O}_2)}{(1 \text{ mol O}_2)} = 109.2 \text{ L O}_2 = \boxed{109 \text{ L O}_2}$$

II. ONE REACTANT IN EXCESS – SINGLE DISPLACEMENT

14 g of lithium is combined with excess K₂SO₄. What is the mass of the new ionic compound that is formed? How do you know that this reaction will take place?

STEP 1: Balance the reaction:



Reaction occurs because Li is higher up than K on the activity series.

STEP 2: Use a molar conversion to convert quantity of reactant given to moles:

$$14 \text{ g Li} * \frac{(1 \text{ mol Li})}{(7 \text{ g Li})} = 2 \text{ mol Li}$$

STEP 3: Use a molar ratio to calculate moles of product that I'm asking about:

$$2 \text{ mol Li} * \frac{(1 \text{ mol Li}_2\text{SO}_4)}{(2 \text{ mol Li})} = 1 \text{ mol Li}_2\text{SO}_4$$

STEP 4: Use a molar conversion to convert moles of product to units that I ask for:

$$1 \text{ mol Li}_2\text{SO}_4 * \frac{(110 \text{ g Li}_2\text{SO}_4)}{(1 \text{ mol Li}_2\text{SO}_4)} = \boxed{110 \text{ Li}_2\text{SO}_4}$$

Section 2 Stoichiometry Answers

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Niculescu,2007-08-07 Creating some links between control feedback and biology modeling communities based on similarities in modeling observing and perceiving alive structures and analyzing interconnections between biological structures and subsystems was the main objective of this volume In this context biology systems need appropriate analysis tools due to their structure and hierarchy complexity and environment interference and we believe that these aspects may generate interesting research topics in control area Indeed several works raising the potential impact of control developments to bring some beginning of answers in the context of biological systems have been published in the recent years The idea of this book was conceived in the context mentioned above with the objective to help in claiming many of the problems for control researchers starting discussions and opening interactive debates between the control and biology communities and finally to alert graduate students to the many interesting ideas at the frontier between control feedback theory and biology [Roadmap to the Regents](#) Sasha Alcott,2003 If Students Need to Know It It s in This Book This book develops the chemistry skills of high school students It builds skills that will help them succeed in school and on the New York Regents Exams Why The Princeton Review We have more than twenty years of experience helping students master the skills needed to excel on standardized tests Each year we help more than 2 million students score higher and earn better grades We Know the New York Regents Exams Our experts at The Princeton Review have analyzed the New York Regents Exams and this book provides the most up to date thoroughly researched practice possible We break down the test into individual skills to familiarize students with the test s structure while increasing their overall skill level We Get Results We know what it takes to succeed in the classroom and on tests This book includes strategies that are proven to improve student performance We provide a breakdown of the skills based on New York standards and objectives hundreds of practice questions organized by skill two complete practice New York Regents Exams in Physical Setting Chemistry **Chemical Principles Study Guide/Solutions Manual** John Krenos,Joseph Potenza,Carl Hoeger,2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

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