20

# ACIDS AND BASES

# PRACTICE PROBLEMS

In your notebook, solve the following problems.

#### SECTION 20.1 DESCRIBING ACIDS AND BASES

Name the following compounds as acids.

a. HNO<sub>2</sub>

b. H<sub>2</sub>SO<sub>4</sub>

c. HF

d. H<sub>2</sub>CO<sub>3</sub>

2. Write the formulas for the following bases.

a. calcium hydroxide

c. aluminum hydroxide

b. ammonium hydroxide

lithium hydroxide

3. Describe some distinctive properties of acids.

Describe some distinctive properties of bases.

#### SECTION 20.2 HYDROGEN IONS AND ACIDITY

- 1. A solution has a hydrogen ion concentration of 1 × 10<sup>-6</sup>M. What is its pH?
- What is the pH of a solution if the [H\*] = 7.2 × 10<sup>-9</sup>M?
- What is the pOH of a solution if the [OH ] = 3.5 × 10<sup>-2</sup>M?
- 4. What is the pOH of a solution that has a pH of 3.4?
- 5. Classify each solution as acidic, basic, or neutral.

a.  $[H^+] = 2.5 \times 10^{-9} M$ 

**d.**  $[H^+] = 1 \times 10^{-7} M$ 

**b.** pOH = 12.0

e. pH = 0.8

c.  $[OH^-] = 9.8 \times 10^{-11} M$ 

6. Calculate the pH of each solution.

a. (H\*) = 1 × 10 5M

c.  $[OH^-] = 2.2 \times 10^{-7}M$ 

**b.**  $[H^*] = 4.4 \times 10^{-13} M$ 

**d.** pOH = 1.4

- 7. Classify the solutions in problem 6 as acidic or basic.
- 8. Why is there a minus sign in the definition of pH?
- 9. A solution has a pOH of 12.4. What is the pH of this solution?
- 10. What is the pH of a solution with [H<sup>-</sup>] = 1 × 10<sup>-3</sup>M?

#### SECTION 20.3 ACID-BASE THEORIES

- Identify the hydrogen ion donor(s) and hydrogen ion acceptor(s) for ionization of H<sub>2</sub>SO<sub>4</sub> in water. Label the conjugate acid-base pairs.
- 2. Identify all of the ions that may be formed when H<sub>3</sub>PO<sub>4</sub> ionizes in water.

# **Section 20 Describing Acids And Bases Answers**

**Neil E. Schore** 

#### **Section 20 Describing Acids And Bases Answers:**

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