

	Covalent Bonds	Ionic Bonds
State at room temperature:	gases, liquids or low-melting solids	Crystalline solids
Polarity:	Low	High
Solubility	Few are soluble in water; Many are soluble in organic liquids	Many are water soluble; Not soluble in organic liquids
Formation:	A covalent bond is formed between two non-metals that have similar electronegativities. Neither atom is "strong" enough to attract electrons from the other. For stabilization, they share their electrons from outer molecular orbit with others	An ionic bond is formed between a metal and a non-metal. Non-metals(-ve ion) are "stronger" than the metal(+ve ion) and can get electrons very easily from the metal. These two opposite ions attract each other and form the ionic bond.
Conductivity:	Do not conduct electricity	Conduct electricity when molten or dissolved in water
Melting point:	Low	High
What is it?:	Covalent bonding is a form of chemical bonding between two non metallic atoms which is characterized by the sharing of pairs of electrons between atoms and other covalent bonds.	Ionic bond, also known as electrovalent bond, is a type of bond formed from the electrostatic attraction between oppositely charged ions in a chemical compound. These kinds of bonds occur mainly between a metallic and a non metallic atom.
Boiling point:	Low	High
Examples:	Methane (CH_4), Hydrochloric acid (HCl)	Sodium chloride (NaCl), Sulfuric Acid (H_2SO_4)
Occurs between:	Two non-metals or a non-metal and a metalloid	One metal and one non-metal

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