Chapter 2: Water

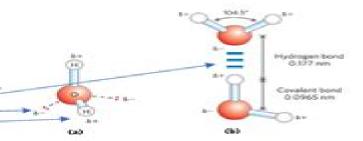
2.1 Water Weak Interactions in Aqueous Systems

Water's Qualities:

- -Unusual properties of water caused by Hydrogen Bonds:
 - 1) High Surface tension beads up
 - 2)High eat capacity
 - 3) High Heat of Vaporization a lot of energy is needed for water to biol
 - 4) Expands as it freezes the solid forms of water (ice) floats instead of sinks like most

Inter/Intramolecular Forces 1-HYDROGEN BONDING

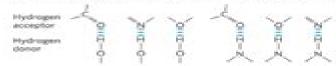
- -Hydrogen bonding occurs when the hydrogen atoms of one molecule is attracted to another molecule that has Nitrogen, Oxygen, and Sulfur (NOS).
- Oxygen is the most electronegative followed by sulfur and Nitrogen.
- -Picture to the right: Water Molecules
 - O is more electronegative than H
 - -This creates a POLAR covalent bond between H and O
 - -The H-Bond occurs between water molecules
 - -O has a partial negative (-) charge
 - -H has a partial positive (+) charge



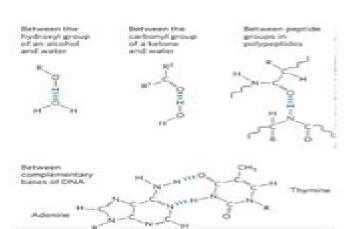
-Polar function groups form H-Bonds with one another and water



-Angle and distance affect the strength of H-Bonds



- -we will see H-Bonds occur between:
 - -Hydroxyl group of alcohol and water
 - -Carbonyl group of a keytone and water
 - -between peptide groups and peptieds
 - -complementary bases of DNA



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David R. Klein

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