

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative	noble gas	cation	electron configuration
positive	anions	anion	polyatomic
compound	oxide	cation	empirical formula

1. An ionic bond is an attraction between oppositely charged _____. **ions**
2. Anionic bond is _____. **negative** change.
3. An ionic compound is formed by losing or gaining _____. **electrons**
4. The _____ noble gases that cannot bond to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both have negative the electron configuration of _____ noble gas.
6. The atoms that make up _____ are ion bonded together by ionic bonds.
7. The _____ method can be used to write the formula for an ionic compound.
8. The _____ empirical formula is a compound denotes the ratio of ions in the compound. If the compound is ionic, write "ion" if it is ionic. Change the subscript word or words to make it ion. Write your answer on the line.

- | | |
|-------------------------|---|
| _____ ionic | 9. Calcium bromide is composed of calcium ion (gaining two electrons). |
| _____ ionic | 10. A cation has a positive charge. |
| _____ ionic | 11. A binary ionic compound contains only one kind of cation and one kind of anion. |
| _____ polyatomic | 12. Substances that consist of more than one atom. |
| _____ cation | 13. The Lewis dot diagram for chlorine, a group 7A element, has six electrons. |
| _____ three (trivalent) | 14. A great deal of energy is produced when an ionic compound is formed. |

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Cangemi



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